THERMOCHEMISTRY CALCULATIONS WORKSHEET 1

Using reaction equation ratios

- How much heat will be released when 6.44 g of Sulfur reacts with excess O₂ according to the following equation?
 2S + 3O₂ → 2SO₃ ΔH² = -791.4kJ
- How much heat will be released when 4.72 g of Carbon reacts with excess O₂ according to the following equation?
 C + O₂ → CO₂ ΔH° = -393.5kJ
- How much heat will be absorbed when 38.2 g of Bromine reacts with excess H₂ according to the following equation?
 H₂ + B₂ → 2HBr ΔH² = 72.80kJ
- How much heat will be released when 1.48 g of Chlorine reacts with excess phosphorus according to the following equation?
 2P + 5Cl₂ → 2PCl₅ ΔH° = -886kJ
- How much heat will be released when 4.77 g of ethanol (C₂H₂OH) reacts with excess O₂ according to the following equation?
 C₂H₂OH + 3O₂ → 2CO₂ + 3H₂O ΔH° = -1366.7kJ
- How much heat will be absorbed when 13.7 g of Nitrogen reacts with excess O₂ according to the following equation?
 N₂ + O₂ → 2NO ΔH° = -180kJ
- How much heat will be released when 11.8 g of Iron reacts with excess O₂ according to the following equation?
 3Fe + 2O₂ → Fe₃O₄ ∆H° = -1120.5kJ
- How much heat will be released when 18.6 g of Hydrogen reacts with excess O₂ according to the following equation?
 2H₂ + O₂ → H₂O ∆H² = -571.6kJ
- How much heat will be transferred when 14.9 g of ammonia reacts with excess O₂ according to the following equation? 4NH₂ + 5O₂ → 4NO + 6H₂O ΔH² = -1170kJ
- How much heat will be transferred when 5.81 g of graphite reacts with excess H₂ according to the following equation?
 6C(graphite) + 3H₂ → C₀H₀ ΔH² = 49.03kJ

Using AH=mCAT

- How many kilojoules of heat energy are required to heat all the aluminum (C_p of Al = .902J/g•°C) in a roll of aluminum foil (500.0 g) from room temperature (25.0 °C) to the temperature of a hot oven (250.0 °C)?
- 12. One way to cool down your cup of coffee is to plunge an ice-cold piece of aluminum into it. Suppose you store a 20.0 g piece of aluminum (C_o of Al = .902J/g•°C) in the refrigerator at 4.40 °C and then drop it into your coffee. The coffee temperature drops from 90.0 °C to 55.0 °C. How many joules of heat energy did the aluminum block absorb? (Ignore the cooling of the cup)
- 13. Suppose you pick up a 16.0-kilogram ball of iron (such as a "shot-put" ball at a track event). The iron (Cp = .451J/g•°C) ball has the same temperature as the atmosphere on a cool day say 16.0 °C. How many kilojoules of heat energy must the iron ball absorb to reach the temperature of your body (37.0°C)?
- 14. The specific heat of silver is 0.24 J/g°C. How much heat in joules must be added to a silver block of mass 86.0 g to raise its temperature by 9.00 °C?
- 15. An 18.7 g sample of platinum metal increases in temperature by 2.30 °C when 5.70 J of heat are added. What is the specific heat of platinum?

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