

c = specific heat ($J/g \cdot ^\circ C$) q = heat energy (J)
 T = temperature ($^\circ C$) m = mass (g) Δ = change in

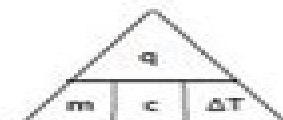
$$q = m \cdot c \cdot \Delta T$$

Name _____ Period _____

Thermal Energy & Heat Unit Study Guide

Part 1. For each of the following scenarios:

- Identify as vaporization, condensation, melting, or freezing
- Write the change in states [example: gas $\rightarrow$ liquid]
- Write whether the process absorbs or releases thermal energy



1. High altitude rain turns to hail.

- _____
- _____

2. Towering cumulonimbus clouds form in the sky.

- _____
- _____

3. A shallow puddle dries up over a couple day span.

- _____
- _____

Part II. Specific Heat Calculations Refer to the equation and symbol legend above. Hint – use the triangle method. Show your work.

- A 110 g sample (m) of tin was heated from $23.2^\circ C$ to $39.8^\circ C$. What was the change in thermal energy? Solve for the heat energy (q). Tin has a specific heat (c) value of $0.2274 J/g \cdot ^\circ C$.

c = _____

m = _____

ΔT = _____

- A scientist heated bundled strands of copper wire from $21.0^\circ C$ to $82.0^\circ C$. Copper has a specific heat (c) capacity of $0.3845 J/g \cdot ^\circ C$. If the change in heat energy (q) was 192 J, what was the mass (m) of the bundled strands of copper wire?

c = _____

ΔT = _____

q = _____

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AW Chickering



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 - The Rise of Digital Reading Thermal Energy And Heat Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Thermal Energy And Heat Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Thermal Energy And Heat Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Thermal Energy And Heat Study Guide
 - Personalized Recommendations
 - Thermal Energy And Heat Study Guide User Reviews and Ratings
 - Thermal Energy And Heat Study Guide and Bestseller Lists
5. Accessing Thermal Energy And Heat Study Guide Free and Paid eBooks
 - Thermal Energy And Heat Study Guide Public Domain eBooks
 - Thermal Energy And Heat Study Guide eBook Subscription Services
 - Thermal Energy And Heat Study Guide Budget-Friendly Options
6. Navigating Thermal Energy And Heat Study Guide eBook Formats

- ePub, PDF, MOBI, and More
- Thermal Energy And Heat Study Guide Compatibility with Devices
- Thermal Energy And Heat Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Thermal Energy And Heat Study Guide
 - Highlighting and Note-Taking Thermal Energy And Heat Study Guide
 - Interactive Elements Thermal Energy And Heat Study Guide
- 8. Staying Engaged with Thermal Energy And Heat Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Thermal Energy And Heat Study Guide
- 9. Balancing eBooks and Physical Books Thermal Energy And Heat Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Thermal Energy And Heat Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Thermal Energy And Heat Study Guide
 - Setting Reading Goals Thermal Energy And Heat Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Thermal Energy And Heat Study Guide
 - Fact-Checking eBook Content of Thermal Energy And Heat Study Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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