

Thermochemistry Review Part 1

Definition Questions: you should be able to use or apply the following terms. Be able to write complete definitions for,

energy	temperature	standard state
potential energy	heat	standard heat of formation
kinetic energy	specific heat capacity	exothermic
enthalpy	endothermic	Law of conservation of Energy
thermal	Hess's Law	isolated, closed, open system

1. The specific heat capacity of diamond is $0.5050 \text{ J/g}^\circ\text{C}$. How much energy is required to heat 25.0 g of diamond from 10.5°C to 15.6°C ?

$$q = m \times c \times \Delta t$$

$$= 25 \text{ g} \times 0.5050 \text{ J/g}^\circ\text{C} \times (15.6 - 10.5^\circ\text{C}) = 64.39 \text{ J}$$

2. A piece of aluminum with a mass of 50 g and an initial temperature of 90°C is placed into 100 mL of water at a temperature of 25°C . The temperature of water rises to 31.3°C . Determine the specific heat capacity of aluminum.

$$Q \text{ lost (Al)} = Q \text{ gained (Water)}$$

$$-(mc\Delta t) = +mc \Delta t$$

$$-(50 \times c \times (31.3 - 90)) = (100 \times 4.18 \times (31.3 - 25))$$

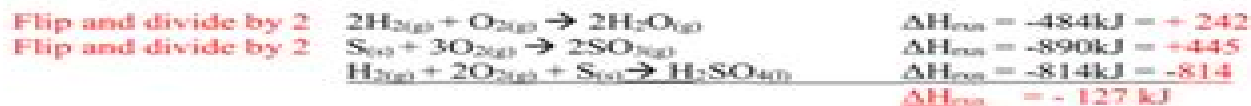
$$-(-2935 c) = 2633.4$$

$$c = 2633.4 / 2935 = 0.897 \text{ J/g}^\circ\text{C}$$

3. How much energy is absorbed by 300 g of methanol, CH_3OH , as it evaporates? Given: The molar heat of vaporization is 35.3 kJ/mol .

$q = ?$	$\Delta H = q/n$	$n \text{ CH}_3\text{OH} = m/\text{MM}$
$m = 300 \text{ g}$	$q = H \times n$	$n = 300 \text{ g} / 32 \text{ g/mol}$
$\Delta H = 35.3 \text{ kJ/mol}$	$q = (35.3 \text{ kJ} \times 9.375 \text{ mol})$	$n = 9.375 \text{ mol}$
	$q = 332.81 \text{ kJ}$	

4. What is the molar enthalpy of the formation of $1 \text{ mol H}_2\text{SO}_{4(l)}$ given the following information?



Thermochem Kinetic Review Answers

Darrell Ebbing



Thermochem Kinetic Review Answers:

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Table of Contents Thermochem Kinetic Review Answers

1. Understanding the eBook Thermochem Kinetic Review Answers
 - The Rise of Digital Reading Thermochem Kinetic Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Thermochem Kinetic Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Thermochem Kinetic Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Thermochem Kinetic Review Answers
 - Personalized Recommendations
 - Thermochem Kinetic Review Answers User Reviews and Ratings
 - Thermochem Kinetic Review Answers and Bestseller Lists
5. Accessing Thermochem Kinetic Review Answers Free and Paid eBooks
 - Thermochem Kinetic Review Answers Public Domain eBooks
 - Thermochem Kinetic Review Answers eBook Subscription Services
 - Thermochem Kinetic Review Answers Budget-Friendly Options
6. Navigating Thermochem Kinetic Review Answers eBook Formats

- ePub, PDF, MOBI, and More
- Thermochem Kinetic Review Answers Compatibility with Devices
- Thermochem Kinetic Review Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Thermochem Kinetic Review Answers
 - Highlighting and Note-Taking Thermochem Kinetic Review Answers
 - Interactive Elements Thermochem Kinetic Review Answers
- 8. Staying Engaged with Thermochem Kinetic Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Thermochem Kinetic Review Answers
- 9. Balancing eBooks and Physical Books Thermochem Kinetic Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Thermochem Kinetic Review Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Thermochem Kinetic Review Answers
 - Setting Reading Goals Thermochem Kinetic Review Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Thermochem Kinetic Review Answers
 - Fact-Checking eBook Content of Thermochem Kinetic Review Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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