Practice Problems Thermochemistry chapter 5

The specific heat of liquid bromine is 0.750 J/g-K. How much heat (J) is required to raise the temperature of 10.0 mL of bromine from 25.00 °C to 35.0°C? The density of liquid bromine is 2.1 g/mL.

Note: the specific heat of 0.750 is bogus. I pulled it out of the air. However, the problem still solves the same way.

 The value of ΔH° for the reaction below is -6535 kJ. _____ kJ of heat are released in the combustion of 16.0 g of C6H6 (1)?

$$2C_6H_6(I) + 15O_2(g) \rightarrow 12CO_2(g) + 6H_2O(I)$$

 Carbon monoxide and oxygen gas react to form carbon dioxide. How much heat is released when 89.5 grams of O₂ (g) reacts with excess CO? ΔH° = -482 kJ.

 The value of ΔH° for the reaction below is -336 kJ. Calculate the heat (kJ) released to the surroundings when 23.0 g of HCl is formed.

$$CH_4(g) + 3Cl_2(g) \rightarrow CHCl_3(1) + 3HCl(g)$$

5. How much heat is released when 29.5 grams of Cl₂ (g) reacts with excess hydrogen?

$$H_2(g) + Cl_2(g) \rightarrow 2HCl(g) \Delta H^0 = -186 \text{ kJ}.$$

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Kaplan Test Prep

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