Answers - Exercises (Thermochemistry)

Q1: Calculate the specific heat capacity of copper given that 204.75 J of energy raises the temperature of 15g of copper from 25°C to 60°C. (Ans: c = 0.39 J/g.K)

Answer:

Q2: A chemist burns 0.8650 g of graphite (a form of carbon) in a new bomb calorimeter, and CO₂ forms. If 393.5 kJ of heat is released per mole of graphite and T increases 2.613 K, what is the capacity of the bomb calorimeter? (Ans: 10.84 kJK)

Answer:

C +
$$O_2$$
 \rightarrow CO_2

Mass of Carbon (Graphite) = 0.8650 g

Atomic mass of Carbon = 12 g/mol

Mol of graphite = mass/atomic mass = 0.8650 g/ 12g mol⁻¹ = 0.072 mol

Heat of combustion, Δ Hcomb = -393.5 kJ/mol

Temperature change, ΔT = 2.613 K

Calorimeter heat capacity, $Ccal$ = ?

$$\Delta$$
Hcomb = $\frac{g_{DD}}{n}$ = $\frac{Ccal \times \Delta T}{n}$

$$393.5$$
 = $\frac{Ccal (2.613)}{0.072}$

Ccal = 10.84 kJ/K

Thermochemistry Exercises Answers

Charles Trapp, Marshall Cady, Carmen Giunta

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Solutions Manual to Accompany Elements of Physical Chemistry C. A. Trapp, M. P. Cady, 2007 The Solutions manual to accompany Elements of Physical Chemistry 4e contains full worked solutions to all end of chapter exercises featured in the book
Thermochemistry Mr. Ajeet Singh, 2018-05-16 Introduces thermodynamic principles related to chemical processes including enthalpy entropy Gibbs free energy and calorimetry
Ions in Solution J Burgess, 1999-10-01 This outline of the principles and chemical interactions in inorganic solution chemistry delivers a course module in an area of considerable complexity Problems with solutions and tutorial hints to test comprehension have been added as a feature to check readers understanding and assist self study Exercises and projects are also provided to help readers deepen and extend their knowledge and understanding Inorganic solution chemistry is treated thoroughly Emphasis is placed upon NMR UV VIS IR Raman spectroscopy X ray diffraction and such topics as acid base behaviour stability constants and kinetics

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