Thermochemistry Practice Problems ANSWER KEY

Calculate the standard enthalpy of reaction for the process

 Calculate the heat evolved in the combustion of ethyl alcohol, ΔH_{comb} C₂H₅ OH₁₈ + 3 O_{2 los} → 2 CO_{2 los} + 3 H₂O₁₈ ΔH = ?

2 C _{graphie} + 3 H _{2 lol} +	1/2 O _{2 tol} → (C₂H₅OH _{ft} ∆H	277.7 kJ
Cgraphes + O2 kg → C	CO2 to	ΔН	393.5 kJ
H _{2 lol} + ½ O _{2 lol} → F	t ₂ O _B	ΔH	285.8 kJ
		AH	= -1366.7 kJ

Calculate the heat required to decompose limestone, CaCO₃, to lime, CaO, and carbon dioxide, with all substances at standard conditions.
 CaCO_{3 N} → CaO_N + CO_{2 N}

ΔH*1 CaCO₃₁₄₀ = -1206.9 kJ/mol ΔH*1 CaO₁₄₀ = -635.1 kJ/mol ΔH*1 CO₃₁₆₀ = -393.5 kJ/mol

 $\Delta H = 178.3 \text{ kJ}$

4. $CH_{Elgl} + 3Cl_{Elgl} \rightarrow CHCl_{Elgl} + 3HCl_{[gl]} \Delta H = ?$

 Cyanamide, CH₂N₂, is a weak acid that is sometimes used as a fertilizer. Calculate the standard enthalpy of formation for cyanamide, given the following standard enthalpies of reaction:

The standard molar enthalpy change is -905.2 kJ for the oxidation of ammonia.
 ⁴ NH_{2 (g)} + 5 O_{2 (g)} → 4 NO (g) + 6 H₂O (g)
 Calculate the standard molar enthalpy of formation for ammonia based on the following standard enthalpies of formation: ΔH*₁ NO (g) = 90.3 kJ/mol and ΔH*₁ H₂O (g) = -241.8 kJ/mol

7. The thermite reaction: 2 Al₁₆₁ + Fe₂O_{3 lot} → Al₂O_{3 lot} + 2 Fe₁₆₁ + 852 kJ

If you begin with 10.0 g of Al and excess Fe₂O₃, how many kJ of heat are evolved at constant pressure?

Thermochemistry Practice Calculations Tesccc Answers

Patrick Vollmar

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