Thermochemistry Practice Problems



Flow of Energy:

1. How many kJ of energy are in a donut that contains 200.0 Calories?

 What is the specific heat of a substance that has a mass of 25.0 g and requires 525.0 calories to raise its temperature by 15.0°C?

$$m = 25.0_{9}$$

 $q = 525.0 \text{ cal}$ $525.0 \text{ cal} = (25.0 \text{ g}) \times (25.0 \text{ c})$
 $\Delta T = 15.0 \text{ c}$ $C = 1.40^{\text{cal}}/\text{g} \cdot \text{c}$

Suppose 100.0 g of H₂O(s) absorbs 1255.0 J of heat. What is the corresponding temperature change?
 The specific heat capacity of H₂O(s) is 2.1 J/g°C. Q = mC Δ T

 How many joules of heat energy are required to raise the temperature of 100.0 g of aluminum by 120.0°C? The specific heat capacity of aluminum is 0.90 J/g°C. Q = MC A T

5. In reference to #4, what is the molar heat capacity of aluminum?

Measuring and Expressing Heat Changes:

6. A student mixed 75.0 mL of water containing 0.75 mol HCl at 25°C with 75.0 mL of water containing 0.75 mol of NaOH at 25°C in a foam cup calorimeter. The temperature of the resulting solution increased to 35°C. How much heat in kilojoules was released by this reaction (Course = 4.18 J/g°C)?

Calculate the amount of heat evolved when 150 g of water are reacted with excess calcium according to the following thermochemical equation:

$$Ca(s) + 2H_2O(l) \rightarrow Ca(OH)_2(s) + H_2(g)$$
 $\Delta H = -65.2 \text{ kJ}$
 $150_3 \text{ HeO} \mid Imol \text{ HeO} \mid -65.2 \text{ kJ} = \left[-270 \text{ kJ} \right] \text{ exothermic}$

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