

THERMOCHEMISTRY CALCULATIONS WORKSHEET 1

Using reaction equation ratios

- How much heat will be released when 6.44 g of Sulfur reacts with excess O_2 according to the following equation?
 $2S + 3O_2 \rightarrow 2SO_3 \quad \Delta H^\circ = -791.4 \text{ kJ}$
- How much heat will be released when 4.72 g of Carbon reacts with excess O_2 according to the following equation?
 $C + O_2 \rightarrow CO_2 \quad \Delta H^\circ = -393.5 \text{ kJ}$
- How much heat will be absorbed when 38.2 g of Bromine reacts with excess H_2 according to the following equation?
 $H_2 + Br_2 \rightarrow 2HBr \quad \Delta H^\circ = 72.80 \text{ kJ}$
- How much heat will be released when 1.48 g of Chlorine reacts with excess phosphorus according to the following equation?
 $2P + 5Cl_2 \rightarrow 2PCl_5 \quad \Delta H^\circ = -886 \text{ kJ}$
- How much heat will be released when 4.77 g of ethanol (C_2H_5OH) reacts with excess O_2 according to the following equation?
 $C_2H_5OH + 3O_2 \rightarrow 2CO_2 + 3H_2O \quad \Delta H^\circ = -1366.7 \text{ kJ}$
- How much heat will be absorbed when 13.7 g of Nitrogen reacts with excess O_2 according to the following equation?
 $N_2 + O_2 \rightarrow 2NO \quad \Delta H^\circ = -180 \text{ kJ}$
- How much heat will be released when 11.8 g of Iron reacts with excess O_2 according to the following equation?
 $3Fe + 2O_2 \rightarrow Fe_3O_4 \quad \Delta H^\circ = -1120.5 \text{ kJ}$
- How much heat will be released when 18.6 g of Hydrogen reacts with excess O_2 according to the following equation?
 $2H_2 + O_2 \rightarrow H_2O \quad \Delta H^\circ = -571.6 \text{ kJ}$
- How much heat will be transferred when 14.9 g of ammonia reacts with excess O_2 according to the following equation?
 $4NH_3 + 5O_2 \rightarrow 4NO + 6H_2O \quad \Delta H^\circ = -1170 \text{ kJ}$
- How much heat will be transferred when 5.81 g of graphite reacts with excess H_2 according to the following equation?
 $6C(\text{graphite}) + 3H_2 \rightarrow C_6H_6 \quad \Delta H^\circ = 49.03 \text{ kJ}$

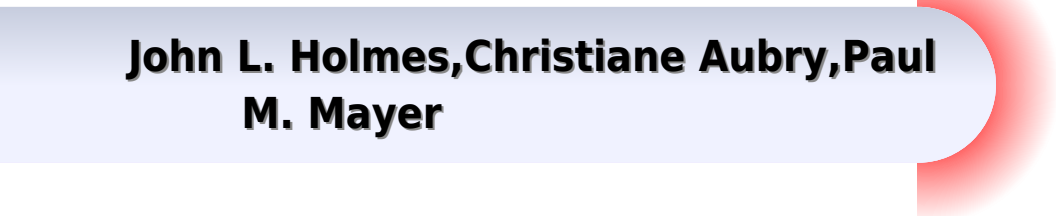
Using $\Delta H = mC\Delta T$

- How many kilojoules of heat energy are required to heat all the aluminum (C_p of Al = $.902 \text{ J/g}^\circ\text{C}$) in a roll of aluminum foil (500.0 g) from room temperature (25.0°C) to the temperature of a hot oven (250.0°C)?
- One way to cool down your cup of coffee is to plunge an ice-cold piece of aluminum into it. Suppose you store a 20.0 g piece of aluminum (C_p of Al = $.902 \text{ J/g}^\circ\text{C}$) in the refrigerator at 4.40°C and then drop it into your coffee. The coffee temperature drops from 90.0°C to 55.0°C . How many joules of heat energy did the aluminum block absorb? (Ignore the cooling of the cup)
- Suppose you pick up a 16.0-kilogram ball of iron (such as a "shot-put" ball at a track event). The iron ($C_p = .451 \text{ J/g}^\circ\text{C}$) ball has the same temperature as the atmosphere on a cool day say 16.0°C . How many kilojoules of heat energy must the iron ball absorb to reach the temperature of your body (37.0°C)?
- The specific heat of silver is $0.24 \text{ J/g}^\circ\text{C}$. How much heat in joules must be added to a silver block of mass 86.0 g to raise its temperature by 9.00°C ?
- An 18.7 g sample of platinum metal increases in temperature by 2.30°C when 5.70 J of heat are added. What is the specific heat of platinum?

Answers: 1) -79.4kJ 2) -155kJ 3) 17.4kJ 4) -3.75kJ 5) -142kJ 6) -88.1kJ 7) -78.7kJ 8) -2660kJ 9) -256kJ 10) 3.96kJ
11) 101kJ 12) 91.2J 13) 152kJ 14) 186J 15) $.132 \text{ J/g}^\circ\text{C}$

Thermochemistry Section Assessment Answers

**John L. Holmes, Christiane Aubry, Paul
M. Mayer**



Thermochemistry Section Assessment Answers:

Thermochemistry Section Assessment Answers Book Review: Unveiling the Magic of Language

In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has become much more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Thermochemistry Section Assessment Answers**," compiled by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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