

Name Key
Specific Heat Worksheet

$C_p = q/m\Delta T$, where q = heat energy, m = mass, and T = temperature

1. A 15.75-g piece of iron absorbs 1086.75 joules of heat energy, and its temperature changes from 25°C to 175°C. Calculate the heat capacity of iron.

$$1086.75 = 15.75g \cdot 150 \cdot C_p$$

$$1086.75 = 2362.5 C_p$$

$$C_p = .46 \text{ J/g}^\circ\text{C}$$

2. How many joules of heat are needed to raise the temperature of 10.0 g of aluminum from 22°C to 55°C, if the specific heat of aluminum is 0.90 J/g°C?

$$q = C_p \cdot m \cdot \Delta T \quad .90(10)(33) =$$

$$297 \text{ J}$$

$$\frac{55}{22} = 33$$

3. To what temperature will a 50.0 g piece of glass raise if it absorbs 5275 joules of heat and its heat capacity is 0.50 J/g°C? The initial temperature of the glass is 20.0°C.

$$5275 = .50(50)\Delta T$$

$$\Delta T = 211$$

$$T_f = \frac{211}{+20} = 231$$

$$231^\circ\text{C}$$

4. Calculate the heat capacity of a piece of wood if 1500.0 g of the wood absorbs 6.75×10^4 joules of heat, and its temperature changes from 32°C to 57°C.

$$67500 = C_p \cdot 1500 \cdot 25^\circ\text{C}$$

$$67500 = 37500 C_p$$

$$1.8 \text{ J/g}^\circ\text{C}$$

5. 100.0 mL of 4°C water is heated until its temperature is 37°C. If the specific heat of water is 4.18 J/g°C, calculate the amount of heat needed to cause this rise in temperature.

$$Q = 100g \cdot 4.18 \cdot 33^\circ\text{C}$$

$$= 13794$$

$$13794 \text{ J}$$

$$13794$$

6. 25.0 g of mercury is heated from 25°C to 155°C, and absorbs 455 joules of heat in the process. Calculate the heat capacity of mercury.

$$455 = C_p \cdot 25g \cdot 130$$

$$455 = 3250 C_p$$

$$.14 = C_p$$

7. What is the Specific heat capacity of silver metal if 55.00 g of the metal absorbs 198 joules of heat and the temperature rises 15.0°C?

$$198 = C_p \cdot 55g \cdot 15^\circ\text{C}$$

$$198 = 825 C_p$$

$$C_p = .24 \text{ J/g}^\circ\text{C}$$

8. What is the change in temperature of 150.0 g of chloroform if it absorbs 1000.0 joules of heat, and the specific heat of chloroform is 0.96 J/g°C?

$$1000 = .96(150g)\Delta T$$

$$1000 = 144 \Delta T$$

$$6.94 = \Delta T$$

$$6.9^\circ\text{C}$$

Answers:

- $C_p = .46 \text{ J/g}^\circ\text{C}$
- $Q = 297 \text{ J}$
- $T_{\text{final}} = 231^\circ\text{C}$
- $C_p = 1.8 \text{ J/g}^\circ\text{C}$
- $Q = 13800 \text{ J}$ 13794
- $C_p = .14 \text{ J/g}^\circ\text{C}$
- $C_p = .24 \text{ J/g}^\circ\text{C}$
- $\Delta T = 6.9^\circ\text{C}$

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Table of Contents Thermochemistry And Specific Heat Problems Canswers

1. Understanding the eBook Thermochemistry And Specific Heat Problems Canswers
 - The Rise of Digital Reading Thermochemistry And Specific Heat Problems Canswers
 - Advantages of eBooks Over Traditional Books
2. Identifying Thermochemistry And Specific Heat Problems Canswers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Thermochemistry And Specific Heat Problems Canswers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Thermochemistry And Specific Heat Problems Canswers
 - Personalized Recommendations

- Thermochemistry And Specific Heat Problems Canswers User Reviews and Ratings
- Thermochemistry And Specific Heat Problems Canswers and Bestseller Lists
- 5. Accessing Thermochemistry And Specific Heat Problems Canswers Free and Paid eBooks
 - Thermochemistry And Specific Heat Problems Canswers Public Domain eBooks
 - Thermochemistry And Specific Heat Problems Canswers eBook Subscription Services
 - Thermochemistry And Specific Heat Problems Canswers Budget-Friendly Options
- 6. Navigating Thermochemistry And Specific Heat Problems Canswers eBook Formats
 - ePub, PDF, MOBI, and More
 - Thermochemistry And Specific Heat Problems Canswers Compatibility with Devices
 - Thermochemistry And Specific Heat Problems Canswers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Thermochemistry And Specific Heat Problems Canswers
 - Highlighting and Note-Taking Thermochemistry And Specific Heat Problems Canswers
 - Interactive Elements Thermochemistry And Specific Heat Problems Canswers
- 8. Staying Engaged with Thermochemistry And Specific Heat Problems Canswers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Thermochemistry And Specific Heat Problems Canswers
- 9. Balancing eBooks and Physical Books Thermochemistry And Specific Heat Problems Canswers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Thermochemistry And Specific Heat Problems Canswers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Thermochemistry And Specific Heat Problems Canswers
 - Setting Reading Goals Thermochemistry And Specific Heat Problems Canswers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Thermochemistry And Specific Heat Problems Canswers
 - Fact-Checking eBook Content of Thermochemistry And Specific Heat Problems Canswers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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