

TEST 2

Q1. What are the units for each of the following if the concentrations are expressed in moles per liter and the time in seconds?

Rate a second-order reaction

Answer 1 Question 1

L/mols

Rate constant for a zero-order reaction

Answer 2 Question 1

L/mols

Rate constant for a first-order reaction

Answer 3 Question 1

1/s

Rate of a first-order reaction

Answer 4 Question 1

mols/s

Q2. The minimum energy required to transform reactants into an activated complex is called _____.

- ☐ a. Kinetic energy
- ☐ b. Heat
- ☒ c. Activation energy
- ☐ d. Potential energy

Q3. A block of copper is heated to 441 °C and then is dropped into a coffee cup calorimeter containing 292 g of water at 25.0 °C. The final temperature of the system is 45.1 °C. What is the mass of the copper block to the nearest gram (1 decimal place)? (Assume all heat is transferred to the water). Write your answer in the box, and do NOT include units.

Specific heat capacity of copper = 0.383 J/ (g°C)

Specific heat capacity of water = 4.18 J/ (g°C)

Answer:

Q4. 2.00M CO and 2.00M H₂O are mixed in a sealed container and the system is left to reach equilibrium. Use the following information to calculate the equilibrium constant, K_c if at equilibrium the concentration of CO₂ is 0.73M



Give your answer in decimal form to 2 significant figures e.g. 0.45

Answer:

Thermodynamics Test Chemistry

Milo D. Koretsky



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Chemical Thermodynamics Maxwell Len McGlashan, 1978 This product is not available separately it is only sold as part of a set There are 750 products in the set and these are all sold as one entity Specialist Periodical Reports provide systematic and detailed review coverage of progress in the major areas of chemical research Written by experts in their specialist fields the series creates a unique service for the active research chemist supplying regular critical in depth accounts of progress in particular areas of chemistry For over 80 years the Royal Society of Chemistry and its predecessor the Chemical Society have been publishing reports charting developments in chemistry which originally took the form of Annual Reports However by 1967 the whole spectrum of chemistry could no longer be contained within one volume and the series Specialist Periodical Reports was born The Annual Reports themselves still existed but were divided into two and subsequently three volumes covering Inorganic Organic and Physical Chemistry For more general coverage of the highlights in chemistry they remain a must Since that time the SPR series has altered according to the fluctuating degree of activity in various fields of chemistry Some titles have remained unchanged while others have altered their emphasis along with their titles some have been combined under a new name whereas others have had to be discontinued **Chemical Thermodynamics** M L

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Thermodynamics for Process Simulation instructs on the behavior of fluids for pure fluids describing the main types of equations of state and their abilities. It discusses the various quantities of interest in process simulation, their correlation and prediction in detail. Chapters look at the important terms for the description of the thermodynamics of mixtures, the most important models and routes for phase equilibrium calculation models which are applicable to a wide variety of non-electrolyte systems, membrane processes, polymer thermodynamics, enthalpy of reaction, chemical equilibria and more. Explains thermodynamic fundamentals used in process simulation with solved examples. Includes new chapters about modern measurement techniques, retrograde condensation and simultaneous description of chemical equilibrium. Comprises numerous solved examples which simplify the understanding of the often complex calculation procedures and discusses advantages and disadvantages of models and procedures. Includes estimation methods for thermophysical properties and phase equilibria, thermodynamics of alternative separation processes. Supplemented with MathCAD sheets and DDBST programs for readers to reproduce the examples. Chemical Thermodynamics for Process Simulation is an ideal resource for those working in the fields of process development, process synthesis or process optimization and an excellent book for students in the engineering sciences.

Physical Chemistry: Thermodynamics Horia Metiu, 2006-02-21. This is a new undergraduate textbook on physical chemistry by Horia Metiu, published as four separate paperback volumes. These four volumes on physical chemistry combine a clear and thorough presentation of the theoretical and mathematical aspects of the subject with examples and applications drawn from current industrial and academic research. By using the computer to solve problems that include actual experimental data, the author is able to cover the subject matter at a practical level. The books closely integrate the theoretical chemistry being taught with industrial and laboratory practice. This approach enables the student to compare theoretical projections with experimental results, thereby providing a realistic grounding for future practicing chemists and engineers. Each volume of Physical Chemistry includes Mathematica and Mathcad Workbooks on CD-ROM. Metiu's four separate volumes: Thermodynamics, Statistical Mechanics, Kinetics, and Quantum Mechanics, offer built-in flexibility by allowing the subject to be covered in any order. These textbooks can be used to teach physical chemistry without a computer, but the experience is enriched substantially for those students who do learn how to read and write Mathematica or Mathcad programs. A TI-89 scientific calculator can be used to solve most of the exercises and problems.

Princeton Review SAT Subject Test Chemistry Prep, 17th Edition The Princeton Review, 2019-12-10. SAT Subject Test Chemistry Prep 17th Edition provides students with a review of all essential content from chemical reactions to kinetics to electron configurations, tons of sample problems and drills, helpful lists of key lab equipment, a cheat sheet of important equations, 3 practice tests and much more. This 17th edition includes a new quick look Study Guide, expanded answer explanations, and access to a new Online Student Tools section with additional college admissions help and info.

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