OMPTER (2) STUDY GUIDE FOR CONTENT MASTERS

## Section 12.3 Limiting Reactants

in your textbook, read about any mactions stop and how to determine the Uniting reaction.

Study the diagram showing a chemical reaction and the eleminal equation that represents the traction. Then complete the table. Show your calculations for questions 25-27 in the space below the table.

$$\Box + 00 + 0 \rightarrow 0 + 00$$
  
 $0,+200 \rightarrow 200$ 

The media masses of O<sub>2</sub>, NO, and NO<sub>2</sub> are 32.00 g/mel, 30.01 g/mel, and 46.01 g/mel, emporatedly.

Amount of O <sub>2</sub>	Annual of NO	Amount of Mily	Similing Number	Amount and Name of Quees Resident
1 molecule	Teolesies	2 molecules	mine	none
Challetales	4 mileciles	4 molecules	NO	2 milesia 0 <sub>y</sub>
I molecules	1 militaire	1. 4 molecules	2. 02	1, 4 molecules NO
3.88 mol -	2.00 mail	4. 2.00 mol	5. none	6. none
438 mol	450 mail	7, 4.00 mail	E. NO	9, 2.00 mol 0 <sub>2</sub>
3.00 mol	7.00 mail	16, 7.00 mol	11, NO	12, 1.50 mol O <sub>2</sub>
1.00 mol	400 mil	13. 2.00 mol	14. O <sub>3</sub>	15. 2.00 mol NO
1.500 mail	0.200 mail	16, 0,200 mol	17, 100	18. 0.400 mol O <sub>2</sub>
name	mag	19. 92.02 g	20. none	21. none
26.00 p.	mmy	22, 46.01 g	20. 02	34. 50.12 g NO
mm <sub>1</sub>	20.00 g	25, 28.7% g	26. 02	37, 1,34 g NO

balanced equation mole ratio = 2 mol NO/1 mol O2

10.00 g.O<sub>2</sub> × 1 mol O<sub>2</sub>/32.00 g.O<sub>2</sub> = 0.3125 mol O<sub>2</sub>

20.00 p.NO × 1 mai NO/30.01 p.NO = 0.6664 mai NO

actual mole ratio = 0.6664 mol NO/0.3125 mol O<sub>2</sub> = 2.132 mol NO/1.000 mol O<sub>2</sub> Secause the actual mole ratio of NO:O<sub>2</sub> is larger than the balanced equation mole ratio of NO:O<sub>3</sub>, there is an excess of NO; O<sub>3</sub> is the limiting reactant.

Mass of NO used = 0.3125 gpai-0; × 2 mol NO<sup>1</sup>1 gpai-0; = 0.6250 mol NO 0.6250 gpai-NO × 30.01 g NO<sup>1</sup>1 gpai-NO = 18.76 g NO

Mass of NO<sub>2</sub> produced = 0.6250 gpsi.WO<sub>2</sub> × 46.01 g NO<sub>2</sub>/1 gpsi.WO<sub>2</sub> = 28.76 g NO<sub>2</sub> Excess NO = 20.00 g NO - 18.76 g NO = 1.24 g NO

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Chemistry Matter and Change + Chapter 12 ... 29.



#### Section 12.4 Percent Yield

In your textlisiol, resel about the pields of products.

Study the diagram and the example problem.

mass of product from experimental measurement

mais of product predicted from assistionietric calculation using

- at mass of reactions.
- 5. 4-dep main to main conversion
  - 1. Sems the belanced chemical equation.
    - Calculate the number of moles of reactions, using moles mass.
    - Calculate the number of moles of product, using the appropriate mole ratio.
    - Calculate the mass of product, using the reciprocal of molar mass.

Example Problem: The following chemical equation represents the production of gallium oriole, a substance used in the manufacturing of some semiconductor devices.

In one experiment, the reaction yielded 7.42 g of the enals from a 7.85 g sample of gallium. Describes the percent yield of this reaction. The molar masses of Ga and Ga<sub>2</sub>O<sub>2</sub> are 46.72 glossl and 167.46 glossl, respectively.

Use the information in the diagram and example problem to evaluate each value or expression below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

- 1. actual right unknown 7.42 g Ga<sub>2</sub>O<sub>3</sub>
- 2. mass of touchast: Total y the connect.
- matter of motion of reactant: 7.00 g/G<sub>0</sub> × 1 mod G<sub>0</sub>
   1 mod G<sub>0</sub>
   2.00 g/G<sub>0</sub> × 1 mod G<sub>0</sub>/68.72 g/G<sub>0</sub>
- 4. number of moles of product 0.000 mol  $G_0 \times \frac{2 \bmod G_0/O_0}{1 \bmod G_0} = \frac{0.100 \bmod G_0 \times 2 \bmod G_0}{2 \bmod G_0/O_0/4 \bmod G_0}$
- 5. theoretical yield. 0:0500 and  $O_{m_0}O_{\gamma} \times \frac{107.44 \pm O_{m_0}O_{\gamma}}{1 \text{ and } O_{m_0}O_{\gamma}}$  connect.
- percent yield: 41.7 g Ga<sub>2</sub>O<sub>3</sub> × 100
   7.42 g Ga<sub>2</sub>O<sub>3</sub>/9.37 g Ga<sub>2</sub>O<sub>3</sub> × 100
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**Charles Secolsky, D. Brian Denison** 

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dimensions of student learning in the 21st century are discussed and a second empirical part in which questions related to how students learning can be measured and how it develops are considered Areas covered include Cultural influences on learning patterns Predicting learning outcomes Student centred learning environments and self directed learning Mathematics learning This indispensable book covers multiple conceptual perspectives on how learning patterns can be described and effects and developments can be measured and will not only be helpful for learning researchers as such but also for educational researchers from the broad domain of educational psychology motivation psychology and instructional sciences who are interested in student motivation self regulated learning effectiveness of innovative learning environments as well as assessment and evaluation of student characteristics and learning process variables Success Sean McWherter, 2017-03-16 Educators often have trouble properly implementing Professional Learning Communities PLCs because they simply don't know how the process is supposed to work By cutting through the fluff and generalities this book provides a clear road map that takes school leaders step by step through the entire PLC process Each chapter addresses a foundational component or protocol necessary for building successful team based learning communities using real life examples to help teachers and leaders understand how to integrate this process and avoid common pitfalls that inhibit implementation Whether you re just starting the PLC process or you re looking to get more out of your PLCs this book will lead you to continued student and teacher growth regardless of current achievement levels socioeconomic status or impending curricular changes Handbook on Measurement, Assessment, and Evaluation in Higher Education Charles Secolsky, D. Brian Denison, 2017-07-31 In this valuable resource well known scholars present a detailed understanding of contemporary theories and practices in the fields of measurement assessment and evaluation with guidance on how to apply these ideas for the benefit of students and institutions Bringing together terminology analytical perspectives and methodological advances this second edition facilitates informed decision making while connecting the latest thinking in these methodological areas with actual practice in higher education This research handbook provides higher education administrators student affairs personnel institutional researchers and faculty with an integrated volume of theory method and application **Teaching at Its Best** Linda B. Nilson, 2010-04-20 Teaching at Its Best This third edition of the best selling handbook offers faculty at all levels an essential toolbox of hundreds of practical teaching techniques formats classroom activities and exercises all of which can be implemented immediately This thoroughly revised edition includes the newest portrait of the Millennial student current research from cognitive psychology a focus on outcomes maps the latest legal options on copyright issues and how to best use new technology including wikis blogs podcasts vodcasts and clickers Entirely new chapters include subjects such as matching teaching methods with learning outcomes inquiry guided learning and using visuals to teach and new sections address Felder and Silverman's Index of Learning Styles SCALE UP classrooms multiple true false test items and much more Praise for the Third Edition of Teaching at Its BestEveryone veterans as well as

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