## **AP Chemistry Final Study Guide**

#### Unit 1: Atomic Structure and Properties (7%-9%)

- At STP, one mole of gas always occupies 22.4 L
- Mass Spectroscopy:
  - Atoms have different isotopes (carbon-12, carbon-14, etc) and this will affect the average atomic mass
  - The most abundant isotope will have the greatest effect
- Coulomb's Law:
  - The greater the charge of the nucleus, the more energy is required to ionize an electron
  - The closer an electron is to the nucleus, the more energy is required to ionize an electron
- · Electromagnetic Radiation:
  - When atoms absorb energy, electrons jump to higher energy levels
  - When electrons drop from higher to lower energy levels, atoms give off electromagnetic radiation energy
- Kinetic Energy of Ionized Electrons:
  - The faster an ejected electron is going, the more kinetic energy it has
  - Electrons originally further away from the nucleus require less energy to eject, so
    it will have more kinetic energy
- · Electron Configuration:
  - When entering the d-block (transition metals), the energy level drops by 1 unless you are using shorthand notation
    - Nickel: 1s<sup>2</sup>2s<sup>2</sup>2p<sup>6</sup>3s<sup>2</sup>3p<sup>6</sup>3d<sup>8</sup>4s<sup>2</sup>(notice the d-block)
    - Nickel (with shorthand): [Ar]4s<sup>2</sup>3d<sup>8</sup>
- Law of Definite Proportions:
  - A pure chemical compound broken down into elements always contains elements of a fixed ratio, independent of where and how it was created
- Law of Multiple Proportions:
  - When elements form compounds, the proportions of the elements in those chemical compounds can be expressed in small whole number ratios
- Aufbau Principle:
  - When building up the electron configuration of an atom, electrons are placed in orbitals, subshells, and shells in order of increasing energy

## **Study Guide For General Chemistry Final**

**AW Chickering** 

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#### **Table of Contents Study Guide For General Chemistry Final**

- 1. Understanding the eBook Study Guide For General Chemistry Final
  - The Rise of Digital Reading Study Guide For General Chemistry Final
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Study Guide For General Chemistry Final
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Study Guide For General Chemistry Final
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Study Guide For General Chemistry Final
  - Personalized Recommendations
  - Study Guide For General Chemistry Final User Reviews and Ratings
  - Study Guide For General Chemistry Final and Bestseller Lists
- 5. Accessing Study Guide For General Chemistry Final Free and Paid eBooks

- Study Guide For General Chemistry Final Public Domain eBooks
- Study Guide For General Chemistry Final eBook Subscription Services
- Study Guide For General Chemistry Final Budget-Friendly Options
- 6. Navigating Study Guide For General Chemistry Final eBook Formats
  - o ePub, PDF, MOBI, and More
  - Study Guide For General Chemistry Final Compatibility with Devices
  - Study Guide For General Chemistry Final Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - o Adjustable Fonts and Text Sizes of Study Guide For General Chemistry Final
  - Highlighting and Note-Taking Study Guide For General Chemistry Final
  - Interactive Elements Study Guide For General Chemistry Final
- 8. Staying Engaged with Study Guide For General Chemistry Final
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Study Guide For General Chemistry Final
- 9. Balancing eBooks and Physical Books Study Guide For General Chemistry Final
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Study Guide For General Chemistry Final
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Study Guide For General Chemistry Final
  - Setting Reading Goals Study Guide For General Chemistry Final
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Study Guide For General Chemistry Final
  - Fact-Checking eBook Content of Study Guide For General Chemistry Final
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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