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web 1. answer a: explanation: the value of ΔH is negative; therefore the reaction is exothermic or heat-releasing. 2. answer a: 1. explanation: the entire reaction was multiplied by 2; therefore ΔH must also be halved. 3. answer c: explanation: option c is incorrect because 566 kJ/mol is the ΔH per mole of O_2 and not for 2 moles.

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web we will revisit this today and go on to see that if we use a special kind of thermochemical reaction called the standard enthalpy of formation we can calculate enthalpies of reactions without having to manipulate a series of individual thermochemical equations for each step.

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