

The Entrance Exam for Prospective Paramedic Students is an EMS-specific exam that was designed to evaluate nationally registered EMTs who are applying or matriculating to paramedic school. The exam has over 130 questions, which range in difficulty level from knowledge recall to problem solving. The exam consists of six topic areas, which were selected by a national committee of Emergency Medical Service educators for the purpose of determining which students are most likely to succeed in a paramedic education program.

After completing this study guide, if you wish to seek additional EMT, anatomy and physiology practice questions, [please visit www.fisdap.net](http://www.fisdap.net) and purchase EMT Study Tools.

Anatomy - EMT level questions ranging from basic to very challenging

1. What are the structures that branch off the trachea into the lower airway?

- A. Alveoli
- B. Bronchi
- C. Cricoid
- D. Arterioles

2. What chamber of the heart is the most muscular?

- A. Left atrium
- B. Right atrium
- C. Left ventricle
- D. Right ventricle

3. What is considered a central pulse?

- A. Radial
- B. Brachial
- C. Dorsalis pedis
- D. Femoral

Physiology - EMT level questions ranging from basic to very challenging

4. What is the amount of air that normally reaches the alveoli in an adult?

- A. 350 mL
- B. 500 mL
- C. 150 L
- D. 750 L

5. Where does the heart receive its nutrition from blood passing through?

- A. Pulmonary artery
- B. Coronary artery
- C. Coronary vein
- D. Aorta

Study Guide For The Sheriff Written Exam

Michal Rosen-Zvi



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Study Guide for the Florida Corrections Officer Certification Exam William G. Doerner, Charles W. Rushing, 2003-01-15
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web the heat of reaction ΔH_{soln} is written after the products in units of kJ mol rxn for example sodium hydroxide dissolves exothermically releasing 44.2 kilojoules per mole dissolved this process is represented as $\text{NaOH (s)} \rightarrow \text{NaOH (aq)}$ $\Delta H_{\text{soln}} = -44.2 \text{ kJ mol rxn}$ write equations to similarly represent the dissolving process for

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web hand warmers are usually placed inside the gloves to keep your hands warm the heat keeping your hands warm is formed as a result of an exothermic reaction which occurs in a small packet hand warmers have 5 main ingredients in them salt carbon polypropylene vermiculite water and iron

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web safety 1 all chemicals in the laboratory are to be considered dangerous do not touch taste or smell any chemicals unless specifically instructed to do so 2 never remove chemicals or other materials from the laboratory area 3 do not eat food drink beverages or chew gum in the laboratory

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web background the hand warmers are filled with sodium acetate click the metal disc and you will notice the liquid changes state and becomes a solid in the process heat is released the sodium acetate is a supersaturated solution

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web oct 9 2021 sodium acetate coin or other small piece of metal plastic bags seal the coin inside a small bag place this bag inside a larger bag containing supersaturated sodium acetate get this by dissolving as much sodium acetate as possible in hot water this chemical hand warmer is reusable

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web the goal of my experiment and research is to evaluate the various number of compounds to be used as hand warmers and locate the ones that can raise the temperature to 20 degrees celsius in 100 g water and costs as little as possible

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web in this lab you are challenged to use chemistry to determine which ionic compound makes the best hand warmer in this experiment you will collect data that will allow you to calculate the heat of solution ΔH_{soln} or the amount of energy released or absorbed in kJ/mol of the substance being dissolved

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web mar 4 2014 contains water the hand warmer is struck in a manner that ruptures the inner pouch releasing the ionic salt into the water of the outer pouch the salt dissolves and the water warms this type of hand warmer tends to produce a more vigorous heat than the dry powder type of hand warmer but does not produce heat for quite as long

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web hand warmers provide a unique and fun way to study the chemistry of crystal formation and heat generation by pressing a button in a pouch which contains a supercooled solution you start a rapid exothermic heat producing crystallization in this science fair project you will determine how the starting temperature affects hand warmer chemistry

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