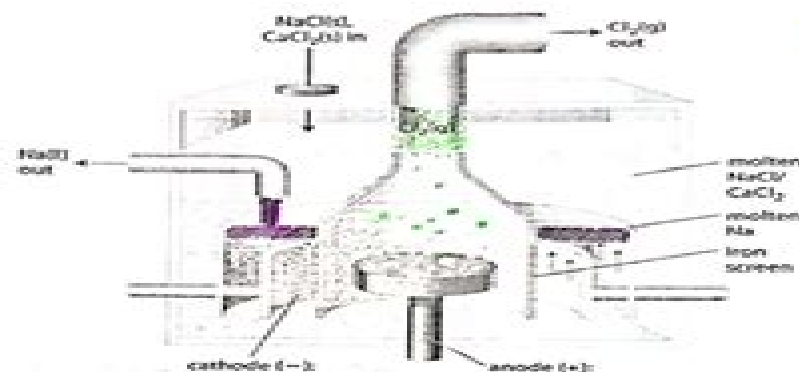


IGCSE – HONORS – ELECTROCHEMISTRY NOTES AND STUDY GUIDE:

Electrolysis of sodium chloride (in the industry)

Industrial electrolysis of molten Sodium Chloride requires extreme temperatures.

- Calcium Chloride is added to the mixture to lower the melting point of Sodium Chloride.
- Molten sodium metal floats in the mixture because of its low density.
- The anode is carbon
- The cathode is iron
- Chlorine gas is collected separately



3.

4. electrolysis of Aluminum oxide (molten)

Electrodes: graphite (carbon)

Electrolyte: molten Alumina (Al_2O_3)

Alumina = Purified Bauxite

Which substance is added to the molten alumina? cryolite

Ions Present: Al^{3+} , O^{2-}

Half equations:

In the Cathode $(Al^{3+} + 3e^- \rightarrow Al) \times 4$

In the Anode $(2O^{2-} \rightarrow O_2 + 4e^-) \times 3$

Overall Equation $4Al^{3+} + 6O^{2-} \rightarrow 4Al + 3O_2$

Observations:

Cathode: The Molten Aluminum can be observed

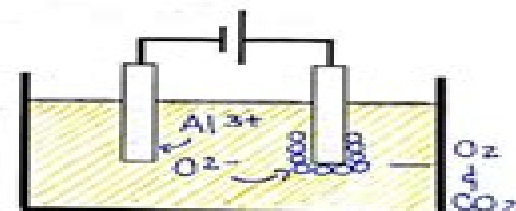
Anode: Bubbles of Oxygen gas will appear

Why is it necessary to change the electrodes periodically?

Because the anodes will burn out. Anode reacts with O_2 to form CO_2

Which is the purpose of the cryolite?

To lower the melting point of the Aluminum Oxide



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Xiang Xie



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Table of Contents Study Guide For Content Mastery Electrochemistry Answers

1. Understanding the eBook Study Guide For Content Mastery Electrochemistry Answers
 - The Rise of Digital Reading Study Guide For Content Mastery Electrochemistry Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide For Content Mastery Electrochemistry Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Study Guide For Content Mastery Electrochemistry Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide For Content Mastery Electrochemistry Answers
 - Personalized Recommendations
 - Study Guide For Content Mastery Electrochemistry Answers User Reviews and Ratings
 - Study Guide For Content Mastery Electrochemistry Answers and Bestseller Lists
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 - Interactive Elements Study Guide For Content Mastery Electrochemistry Answers
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- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Study Guide For Content Mastery Electrochemistry Answers
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- Dealing with Digital Eye Strain
 - Minimizing Distractions
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- Setting Reading Goals Study Guide For Content Mastery Electrochemistry Answers
 - Carving Out Dedicated Reading Time
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- Fact-Checking eBook Content of Study Guide For Content Mastery Electrochemistry Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

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- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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