

Basic Chemistry Study Guide 2

1. Hydrogen bonds

1. Form between two polar molecules **covalent bonds**
2. A hydrogen's δ^+ region is attracted to a δ^- region in another molecule.
3. Form between water molecules (Where are the δ^+ and δ^- regions in a water molecule? **hydrogen bonds, at the top of the symbol**)
4. Each hydrogen bond is very weak (can be easily formed and broken).
5. There are many hydrogen bonds in water. Together, these bonds give water its unique properties that are essential for life
6. One of these properties is the ability to be a solvent for other polar molecules.
7. Hydrogen bonds are depicted in figures as dotted or dashed lines between two atoms.

2. Solutions, solvents, and solutes

1. Water is the most abundant molecule in living organisms, and 60% of human body weight is water.
2. Water is located inside and outside of our cells.
3. Body fluids are solutions
4. Define the terms **solution, solvent, and solute**. **Solvent: a substance capable of dissolving other polar molecules and ionic bonds, solutes: molecules dissolved in the solvent, solution: homogeneous mixture**

3. How substances dissolve in water

1. When a solute dissolves, it becomes surrounded by water molecules.
2. For a solute to dissolve, it must attract water molecules.
3. What kinds of substances attract water molecules? **Sugars, nucleic acids, and amino acids**
4. When a solute dissolves it forms hydrogen bonds with water molecules.
5. Define the terms **ion, cation, and anion**. **ion: atom with a charge, cation: positively charged ion, anions: negatively charged ion**
6. Define the terms **hydrophilic** and **hydrophobic** **Hydrophilic: attracted to water, hydrophobic: afraid of water**
7. What types of molecules are hydrophilic and what types are hydrophobic? **Nonpolar hydrophobic, polar hydrophilic**

Study Guide Of Foundations Of College Chemistry

Morris Hein, Susan Arena



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Table of Contents Study Guide Of Foundations Of College Chemistry

1. Understanding the eBook Study Guide Of Foundations Of College Chemistry
 - The Rise of Digital Reading Study Guide Of Foundations Of College Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide Of Foundations Of College Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Study Guide Of Foundations Of College Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide Of Foundations Of College Chemistry
 - Personalized Recommendations
 - Study Guide Of Foundations Of College Chemistry User Reviews and Ratings
 - Study Guide Of Foundations Of College Chemistry and Bestseller Lists
5. Accessing Study Guide Of Foundations Of College Chemistry Free and Paid eBooks
 - Study Guide Of Foundations Of College Chemistry Public Domain eBooks
 - Study Guide Of Foundations Of College Chemistry eBook Subscription Services
 - Study Guide Of Foundations Of College Chemistry Budget-Friendly Options

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 - Study Guide Of Foundations Of College Chemistry Enhanced eBook Features
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 - Adjustable Fonts and Text Sizes of Study Guide Of Foundations Of College Chemistry
 - Highlighting and Note-Taking Study Guide Of Foundations Of College Chemistry
 - Interactive Elements Study Guide Of Foundations Of College Chemistry
8. Staying Engaged with Study Guide Of Foundations Of College Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Study Guide Of Foundations Of College Chemistry
9. Balancing eBooks and Physical Books Study Guide Of Foundations Of College Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Study Guide Of Foundations Of College Chemistry
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Study Guide Of Foundations Of College Chemistry
 - Setting Reading Goals Study Guide Of Foundations Of College Chemistry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Study Guide Of Foundations Of College Chemistry
 - Fact-Checking eBook Content of Study Guide Of Foundations Of College Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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 - Integration of Multimedia Elements

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