

How To Find Theoretical Yield

Theoretical yield is the mass of product that the limiting reactant produces. All we need to do is to convert moles of the product that the limiting reactant gives to grams.

Example: If 10 grams of Zn reacted with 20 grams of HCl, what is the theoretical yield of hydrogen gas?



Transformation
TUTORING

$$\text{Zn: } 10\text{g Zn} \times \frac{1\text{mol Zn}}{65.38\text{g Zn}} \times \frac{1\text{mol H}_2}{1\text{mol Zn}} = 0.153\text{mol H}_2$$

limiting

SMALLER

$$\text{HCl: } 20\text{g HCl} \times \frac{1\text{mol HCl}}{36.458\text{g HCl}} \times \frac{1\text{mol H}_2}{2\text{mol HCl}} = 0.274\text{mol H}_2$$

$$0.153\text{mol H}_2 \times \frac{2.016\text{g}}{1\text{mol H}_2} = 0.308\text{g H}_2$$

theoretical yield of hydrogen gas

Theoretical Yield And Limiting Reagents Answer Key

Rachel Sandford



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