

How To Find Theoretical Yield

Theoretical yield is the mass of product that the limiting reactant produces. All we need to do is to convert moles of the product that the limiting reactant gives to grams.

Example: If 10 grams of Zn reacted with 20 grams of HCl, what is the theoretical yield of hydrogen gas?



Transformation
TUTORING

$$\text{Zn: } 10\text{g Zn} \times \frac{1\text{mol Zn}}{65.38\text{g Zn}} \times \frac{1\text{mol H}_2}{1\text{mol Zn}} = 0.153\text{mol H}_2$$

limiting

SMALLER

$$\text{HCl: } 20\text{g HCl} \times \frac{1\text{mol HCl}}{36.458\text{g HCl}} \times \frac{1\text{mol H}_2}{2\text{mol HCl}} = 0.274\text{mol H}_2$$

$$0.153\text{mol H}_2 \times \frac{2.016\text{g}}{1\text{mol H}_2} = 0.308\text{g H}_2$$

theoretical yield of hydrogen gas

Theoretical Yield Limiting Reagents Answer Key

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Introduction to General, Organic, and Biochemistry Morris Hein, Scott Pattison, Susan Arena, Leo R. Best, 2014-01-15 The most comprehensive book available on the subject Introduction to General Organic and Biochemistry 11th Edition continues its tradition of fostering the development of problem solving skills featuring numerous examples and coverage of current applications Skillfully anticipating areas of difficulty and pacing the material accordingly this readable work provides clear and logical explanations of chemical concepts as well as the right mix of general chemistry organic chemistry and biochemistry An emphasis on real world topics lets readers clearly see how the chemistry will apply to their career

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web jun 2 2020 the reactant that is consumed first and limits the amount of product s that can be obtained is the limiting reactant to identify the limiting reactant calculate the number of moles of each reactant present and compare this ratio to the mole ratio of the reactants in the balanced chemical equation

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