

8. When you make ice cubes, the entropy of water ? [AIIMS-2006]
 (1) Does not change
 (2) Increases
 (3) Decreases
 (4) May either increase or decrease depending on the process used
9. For a phase change $\text{H}_2\text{O}(l) \rightleftharpoons \text{H}_2\text{O}(s)$ [AIIMS-2006]
 (1) $\Delta G = 0$ (2) $\Delta S = 0$ (3) $\Delta H = 0$ (4) $\Delta U = 0$
10. For a spontaneous process the correct statement is - [AIIMS-2006]
 (1) Entropy of the system always increase
 (2) Free energy of the system always increases
 (3) Total entropy change is always negative
 (4) Total entropy change is always positive
11. The enthalpy change (ΔH) for the reaction, $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \longrightarrow 2\text{NH}_3(\text{g})$ is -92.38 kJ at 298 K . The internal energy change ΔU at 298 K is - [AIIMS-2006]
 (1) -92.38 kJ (2) -87.42 kJ (3) -97.34 kJ (4) -89.9 kJ
12. Consider the following reactions : [AIPMT-2007]
 (a) $\text{H}^+_{(\text{aq})} + \text{OH}^-_{(\text{aq})} = \text{H}_2\text{O}_{(\text{l})}$, $\Delta H = -X_1 \text{ kJ mol}^{-1}$
 (b) $\text{H}_{2(\text{g})} + \frac{1}{2} \text{O}_{2(\text{g})} = \text{H}_2\text{O}_{(\text{l})}$, $\Delta H = -X_2 \text{ kJ mol}^{-1}$
 (c) $\text{CO}_{2(\text{g})} + \text{H}_{2(\text{g})} = \text{CO}_{(\text{g})} + \text{H}_2\text{O}_{(\text{l})} - X_3 \text{ kJ mol}^{-1}$
 (d) $\text{C}_2\text{H}_{2(\text{g})} + \frac{5}{2} \text{O}_{2(\text{g})} = 2\text{CO}_{2(\text{g})} + \text{H}_2\text{O}_{(\text{l})} + X_4 \text{ kJ mol}^{-1}$
 Enthalpy of formation of $\text{H}_2\text{O}_{(\text{l})}$ is :
 (1) $+X_1 \text{ kJ mol}^{-1}$ (2) $-X_2 \text{ kJ mol}^{-1}$ (3) $+X_3 \text{ kJ mol}^{-1}$ (4) $-X_4 \text{ kJ mol}^{-1}$
13. Given that bond energies of H-H and Cl-Cl are 430 kJ mol^{-1} and 240 kJ mol^{-1} respectively and $\Delta_f H$ for HCl is -90 kJ mol^{-1} . Bond enthalpy of HCl is : [AIPMT-2007]
 (1) 245 kJ mol^{-1} (2) 290 kJ mol^{-1} (3) 380 kJ mol^{-1} (4) 425 kJ mol^{-1}
14. Bond dissociation enthalpy of H_2 , Cl_2 and HCl are 434 , 242 and 431 kJ mol^{-1} respectively. Enthalpy of formation of HCl is :- [AIPMT-2008]
 (1) -93 kJ mol^{-1} (2) 245 kJ mol^{-1} (3) 93 kJ mol^{-1} (4) -245 kJ mol^{-1}
15. For the gas phase reaction, [AIPMT-2008]
 $\text{PCl}_5(\text{g}) \rightleftharpoons \text{PCl}_3(\text{g}) + \text{Cl}_2(\text{g})$
 Which of the following conditions are correct ?
 (1) $\Delta H < 0$ and $\Delta S < 0$ (2) $\Delta H > 0$ and $\Delta S < 0$
 (3) $\Delta H = 0$ and $\Delta S < 0$ (4) $\Delta H > 0$ and $\Delta S > 0$

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