

Thermochemistry Practice Problems

KEY

Flow of Energy:

1. How many kJ of energy are in a donut that contains 200.0 Calories?

$$\frac{200 \text{ Cal}}{1 \text{ Cal}} \times \frac{4.184 \text{ J}}{1 \text{ Cal}} = \boxed{836.8 \text{ kJ}}$$

2. What is the specific heat of a substance that has a mass of 25.0 g and requires 525.0 calories to raise its temperature by 15.0°C?

$$\begin{aligned} m &= 25.0 \text{ g} \\ q &= 525.0 \text{ cal} \\ \Delta T &= 15.0^\circ\text{C} \end{aligned}$$

$$q = mc\Delta T$$

$$525.0 \text{ cal} = (25.0 \text{ g})(c)(15.0^\circ\text{C})$$

$$\boxed{c = 1.40 \text{ cal/g}^\circ\text{C}}$$

3. Suppose 100.0 g of H₂O(s) absorbs 1255.0 J of heat. What is the corresponding temperature change? The specific heat capacity of H₂O(s) is 2.1 J/g°C.

$$\begin{aligned} q &= 1255.0 \text{ J} \\ m &= 100.0 \text{ g} \\ c &= 2.1 \text{ J/g}^\circ\text{C} \end{aligned}$$

$$q = mc\Delta T$$

$$1255.0 \text{ J} = (100.0 \text{ g})(2.1 \text{ J/g}^\circ\text{C})(\Delta T)$$

$$\boxed{\Delta T = 6.0^\circ\text{C}}$$

4. How many joules of heat energy are required to raise the temperature of 100.0 g of aluminum by 120.0°C? The specific heat capacity of aluminum is 0.90 J/g°C.

$$\begin{aligned} m &= 100.0 \text{ g} \\ c &= 0.90 \text{ J/g}^\circ\text{C} \\ \Delta T &= 120.0^\circ\text{C} \end{aligned}$$

$$q = (100.0 \text{ g})(0.90 \text{ J/g}^\circ\text{C})(120.0^\circ\text{C})$$

$$\boxed{q = 11,000 \text{ J}}$$

5. In reference to #4, what is the molar heat capacity of aluminum?

$$(0.90 \text{ J/g}^\circ\text{C})(26.98 \text{ g/mol}) = \boxed{24 \text{ J/mol}^\circ\text{C}}$$

Measuring and Expressing Heat Changes:

6. A student mixed 75.0 mL of water containing 0.75 mol HCl at 25°C with 75.0 mL of water containing 0.75 mol of NaOH at 25°C in a foam cup calorimeter. The temperature of the resulting solution increased to 35°C. How much heat in kilojoules was released by this reaction ($C_{\text{water}} = 4.18 \text{ J/g}^\circ\text{C}$)?

$$q = mc\Delta T$$

$$m = 150.0 \text{ g}$$

$$c = 4.184 \text{ J/g}^\circ\text{C}$$

$$\Delta T = 10^\circ\text{C}$$

$$q = (150.0 \text{ g})(4.184 \text{ J/g}^\circ\text{C})(10^\circ\text{C})$$

$$\boxed{q = 6300 \text{ J}}$$

7. Calculate the amount of heat evolved when 150 g of water are reacted with excess calcium according to the following thermochemical equation:



$$\frac{150 \text{ g H}_2\text{O}}{18.02 \text{ g H}_2\text{O}} \times \frac{1 \text{ mol H}_2\text{O}}{2 \text{ mol H}_2\text{O}} \times \frac{-65.2 \text{ kJ}}{1} = \boxed{-270 \text{ kJ}; \text{ exothermic}}$$

Thermochemistry Practice Problems And Answers

JL Elias



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Table of Contents Thermochemistry Practice Problems And Answers

1. Understanding the eBook Thermochemistry Practice Problems And Answers
 - The Rise of Digital Reading Thermochemistry Practice Problems And Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Thermochemistry Practice Problems And Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Thermochemistry Practice Problems And Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Thermochemistry Practice Problems And Answers

- Personalized Recommendations
 - Thermochemistry Practice Problems And Answers User Reviews and Ratings
 - Thermochemistry Practice Problems And Answers and Bestseller Lists
5. Accessing Thermochemistry Practice Problems And Answers Free and Paid eBooks
 - Thermochemistry Practice Problems And Answers Public Domain eBooks
 - Thermochemistry Practice Problems And Answers eBook Subscription Services
 - Thermochemistry Practice Problems And Answers Budget-Friendly Options
 6. Navigating Thermochemistry Practice Problems And Answers eBook Formats
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 - Thermochemistry Practice Problems And Answers Compatibility with Devices
 - Thermochemistry Practice Problems And Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Thermochemistry Practice Problems And Answers
 - Highlighting and Note-Taking Thermochemistry Practice Problems And Answers
 - Interactive Elements Thermochemistry Practice Problems And Answers
 8. Staying Engaged with Thermochemistry Practice Problems And Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Thermochemistry Practice Problems And Answers
 9. Balancing eBooks and Physical Books Thermochemistry Practice Problems And Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Thermochemistry Practice Problems And Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Thermochemistry Practice Problems And Answers
 - Setting Reading Goals Thermochemistry Practice Problems And Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Thermochemistry Practice Problems And Answers

- Fact-Checking eBook Content of Thermochemistry Practice Problems And Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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