

Chemistry 30 Thermochemistry Quiz #2 Practice

Use the following information to answer the next question:

The combustion of ethanol ($\text{C}_2\text{H}_5\text{OH}$) is exothermic. Write the equation representing the following thermochemical equation:



Equation:

- | | |
|--|--|
| <p>A. $\text{C}_2\text{H}_5\text{OH}(l) + 3.5 \text{O}_2(g) \rightarrow 2 \text{CO}_2(g) + 1.5 \text{H}_2\text{O}(g) \quad \Delta H = -1366.8 \text{ kJ}$</p> | <p>A. $\Delta H = -1366.8 \text{ kJ}$</p> |
| <p>B. $\text{C}_2\text{H}_5\text{OH}(l) + \frac{1}{2} \text{O}_2(g) \rightarrow 2 \text{CO}_2(g) + 3 \text{H}_2\text{O}(g) \quad \Delta H = -1366.8 \text{ kJ}$</p> | <p>B. $\Delta H = -1366.8 \text{ kJ}$</p> |

4. The enthalpy change for the combustion of ethanol as shown in the thermochemical equation is:

- A. -1366.8 kJ
 B. -683.4 kJ
 C. -1366.8 kJ
 D. -683.4 kJ

Use the following information to answer the next question:

When ethanally, hydrogen combustion, as represented by the following equation:



5. The maximum heat enthalpy information with the reaction is

- A. -1366.8 kJ/mol
 B. -683.4 kJ/mol
 C. -1366.8 kJ/mol
 D. -683.4 kJ/mol

Thermochemistry Question Answer Guide

Raffaella Di Napoli



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