BITTHEY.			

Chemistry 30

Thermochemistry Review

- 1. Which of the following statements is true for an endothermic reaction?
 - A. ΔH_e for the reaction is negative.
 - B. Energy is a product in the chemical equation.
 - C. The temperature of the surroundings increases.
 - D. The potential energy of the products is greater than the potential energy of the reactants.

Use the following information to answer the next question.

$$4 Q(s) + Z_2(g) + 100 kJ \rightarrow 2 Q_2Z(g)$$

- The ΔH_r value for the reverse reaction would be
 - A. +100 kJ
 - B. +50.0 kJ
 - C. -50.0 kJ
 - D. -100 kJ

Numerical Response

Benzene is a gasoline additive. The heat of formation for 1.00 mol of C₆H₆(1) is kJ.

(Record your three-digit answer in the numerical-response section on the answer sheet)

Numerical Response

An outdoor mercury thermometer contains 1.02 g of mercury. On a winter day, the thermometer records a change in temperature from −38.0°C to −13.0°C. The mercury absorbed ______ J of energy. The heat capacity of mercury is 0.140 J/g°C.

(Record your three-digit answer in the numerical-response section on the answer sheet)

Thermochemistry Review Sheet

Lingjun Ying

Thermochemistry Review Sheet:

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Table of Contents Thermochemistry Review Sheet

- 1. Understanding the eBook Thermochemistry Review Sheet
 - The Rise of Digital Reading Thermochemistry Review Sheet
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Thermochemistry Review Sheet
 - Exploring Different Genres
 - o Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Thermochemistry Review Sheet
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Thermochemistry Review Sheet
 - Personalized Recommendations
 - Thermochemistry Review Sheet User Reviews and Ratings
 - Thermochemistry Review Sheet and Bestseller Lists

- 5. Accessing Thermochemistry Review Sheet Free and Paid eBooks
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 - Highlighting and Note-Taking Thermochemistry Review Sheet
 - Interactive Elements Thermochemistry Review Sheet
- 8. Staying Engaged with Thermochemistry Review Sheet
 - o Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Thermochemistry Review Sheet
- 9. Balancing eBooks and Physical Books Thermochemistry Review Sheet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Thermochemistry Review Sheet
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Thermochemistry Review Sheet
 - Setting Reading Goals Thermochemistry Review Sheet
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Thermochemistry Review Sheet
 - Fact-Checking eBook Content of Thermochemistry Review Sheet
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
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