Thermochemistry Formula Sheet:

1 st Law of Thermodynamics:	Internal Energy:
Endothermic: $q=+\ell$ Exothermic: $q=-\ell$	$\Delta U = q + W$
Work Done By/On a Gas:	$W = -P \Delta V$
Work done by the system: W=−i Work done on the system: W=+i	1 L +1 otm = 101.3 Joules
Specific Heat Capacity: J/(g&*C)&	$C_{s} = \frac{q}{m \Delta T}$
Molar Heat Capacity: J/(mol °C)	$C_{\omega} = \frac{q}{n\Delta T}$
Heat Capacity: J/℃	$C_{N} = \frac{q}{\Delta T}$
Calorimetry - Temperature Change:	$q = mc \Delta T$ $C_{H20} = 4.184 \frac{J}{\alpha^{9}C} C_{tot} = 2.09 \frac{J}{\alpha^{9}C} C_{3res} = 2.03 \frac{J}{\alpha^{9}C}$
Calorimetry - Phase Change:	$q=m\Delta H q=n\Delta H$
	H2O: ΔH_{fes} =334 $J/g \Delta H_{fes}$ =6.01 kJ/mol H2O: ΔH_{sap} =2260 $J/g \Delta H_{sap}$ =40.7 kJ/mol
Final Temperature of a Mixture:	$-q_A = q_B$
T▲→ Initial Temp. of Substance A	$-m_{A}C_{A}[T_{F}-T_{A}]=m_{0}C_{6}(T_{F}-T_{6})$
T _B → Initial Temp. of Substance B	Note: This equation will work if there are no phase changes.
Coffee Cup Calorimeter:	$\Delta H^{\circ}_{cos} = \frac{q_{cos}}{n} q_{cos} = -q_{eq,o}$
Enthalpy Change:	

Thermochemistry Cp Test Sheet

Padhraic Smyth

Thermochemistry Cp Test Sheet:

Reviewing Thermochemistry Cp Test Sheet: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is actually astonishing. Within the pages of "**Thermochemistry Cp Test Sheet**," an enthralling opus penned by a highly acclaimed wordsmith, readers set about an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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