

Thermochemistry Worksheets

4 Worksheets | Specific Heat | Heat of Reactions |
Phase Changes | Enthalpy Calculations

Thermochemistry Worksheet 1

Conversions: (1 calorie = 4.18 Joules)

1. How many calories are equivalent to 95.4 Joules?
2. Convert 42.7 calories to Joules.
3. Convert 828 calories to Joules.
4. Convert 912.3 Joules to calories.
5. How many calories are in 312.8 Joules?

Heat Calculations:

Heat = mass $\times$ specific heat $\times$ change in temperature

6. Calculate the heat energy when a stone cools from a heat of 0.218 cal/g°C and a mass of 15.0 g.

7. Determine the heat lost by a metal sample.

Name _____

Thermochemistry Enthalpy Worksheet

Directions: Solve the following enthalpy-related problems using the given thermodynamic data.

g of CaCO_3 forms from the reaction of CaO(s) and $\text{CO}_2\text{(g)}$. $\Delta H_{\text{rxn}} = -281.2 \text{ kJ}$

g of methane (CH_4) combusts completely in oxygen. $\Delta H_{\text{rxn}} = -2890.2 \text{ kJ}$

50 moles of Cu_2O based on the equation:
 $\text{Cu}_2\text{O} \rightarrow 2\text{Cu} + \text{O}_2$

$\Delta H_{\text{rxn}} = -446.1 \text{ kJ/mol}$, how much heat is released?
(1700)

Name _____

Phase Change Enthalpy Worksheet

Directions: Solve the following problems related to enthalpy changes during phase transitions of water. Use the following values:

- Specific heat of water: $4.184 \text{ J/g}^\circ\text{C}$
- Heat of fusion of water: 334 J/g
- Heat of vaporization of water: 2260 J/g

1. A 25 g sample of ice is at -10°C . How much heat is required to raise its temperature to 0°C before it melts?
2. How much heat is needed to melt a 10 g ice cube at 0°C into liquid water?

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Ibrahim Sikder



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1. Understanding the eBook Thermochemistry Heat And Chemical Change Key
 - The Rise of Digital Reading Thermochemistry Heat And Chemical Change Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Thermochemistry Heat And Chemical Change Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Thermochemistry Heat And Chemical Change Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Thermochemistry Heat And Chemical Change Key
 - Personalized Recommendations
 - Thermochemistry Heat And Chemical Change Key User Reviews and Ratings
 - Thermochemistry Heat And Chemical Change Key and Bestseller Lists

5. Accessing Thermochemistry Heat And Chemical Change Key Free and Paid eBooks
 - Thermochemistry Heat And Chemical Change Key Public Domain eBooks
 - Thermochemistry Heat And Chemical Change Key eBook Subscription Services
 - Thermochemistry Heat And Chemical Change Key Budget-Friendly Options
6. Navigating Thermochemistry Heat And Chemical Change Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Thermochemistry Heat And Chemical Change Key Compatibility with Devices
 - Thermochemistry Heat And Chemical Change Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Thermochemistry Heat And Chemical Change Key
 - Highlighting and Note-Taking Thermochemistry Heat And Chemical Change Key
 - Interactive Elements Thermochemistry Heat And Chemical Change Key
8. Staying Engaged with Thermochemistry Heat And Chemical Change Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Thermochemistry Heat And Chemical Change Key
9. Balancing eBooks and Physical Books Thermochemistry Heat And Chemical Change Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Thermochemistry Heat And Chemical Change Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Thermochemistry Heat And Chemical Change Key
 - Setting Reading Goals Thermochemistry Heat And Chemical Change Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Thermochemistry Heat And Chemical Change Key
 - Fact-Checking eBook Content of Thermochemistry Heat And Chemical Change Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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