

Thermochemistry Problems

The First Law

1. The complete combustion of acetic acid, $\text{HC}_2\text{H}_3\text{O}_2$, to form water, H_2O , and CO_2 , at constant pressure releases 871.7 kJ of heat per mol of acetic acid. Write a balanced thermochemical equation for this reaction. Draw an enthalpy diagram for the reaction.
2. Consider the following reaction, which occurs at standard state:
$$2\text{Cl}_{(g)} \rightarrow \text{Cl}_{2(g)} \quad \Delta H^\circ = -243.4 \text{ kJ}$$
 - a. Which has the higher, more positive, enthalpy under these conditions, $2\text{Cl}_{(g)}$ or $\text{Cl}_{2(g)}$?
 - b. Consider the following reaction: $2 \text{Mg}_{(s)} + \text{O}_{2(g)} \rightarrow 2 \text{MgO}_{(s)}$ $\Delta H = -1204 \text{ kJ}$.
 - c. Is the reaction exothermic or endothermic?
 - d. Calculate the amount of heat transferred when 2.4 g of $\text{Mg}_{(s)}$ reacts at constant pressure. How many grams of MgO are produced during an enthalpy change of 96.0 kJ?
 - e. How many kilo joules of heat are absorbed when 7.50 g of $\text{MgO}_{(s)}$ are decomposed into $\text{Mg}_{(s)}$ and $\text{O}_{2(g)}$ at constant pressure?

Enthalpy

Calorimetry Problems

3. Determine the specific heat of a sample of Cu from the fact that 64.0J are needed to raise the temperature of 15.0 g of Cu metal from 22.0°C to 33.0°C. [0.388 J/g°C]
4. The specific heat of a sample of copper is 0.385 J/g-K. How many joules of heat are necessary to raise the temperature of a 1.42 kg block of copper from 25.0 °C to 88.5°C?
5. A 50.0-g sample of water at 100.00°C was placed in an insulated cup. Then 25.3-g of zinc at 25.00°C was added to the water. The temperature of the water dropped to 96.68°C. What is the specific heat of the zinc?[0.388 J/g°C]
6. When a 6.50 g sample of solid NaOH dissolves in 100.0g of water in a coffee-cup calorimeter, the temperature rises from 21.6°C to 37.8°C. Calculate the ΔH (in kJ/mol NaOH) for the solution process. (Assume the specific heat of the solution formed is 4.18 J/g-K)
$$\text{NaOH}_{(s)} \rightarrow \text{Na}^+(\text{aq}) + \text{OH}^-(\text{aq})$$
7. A house is designed to have passive solar energy features, Brickwork is to be incorporated into the interior of the house to act as a heat absorber. Each brick

Thermochemistry Problems And Solutions

SJ Ball



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