

THERMODYNAMICS TEST REVIEW ANSWERS

Specific Heat Problems

1. How many J would it take to raise the temperature of 200 grams of water from 5 degree Celsius C to 85 degree Celsius?

Answer: $Q = (200 \text{ g})(4.18)(80 \text{ }^{\circ}\text{C}) = 66880 \text{ J}$

2. How many J would problem number 1 be if it was aluminum instead of water?

Answer: $Q = (200 \text{ g})(0.897)(80 \text{ }^{\circ}\text{C}) = 14352 \text{ J}$

3. How many grams of copper could be heated from 20 °C to 75 °C if 1200J are applied to it?

Answer: $1200 \text{ J} = (x)(0.385)(55 \text{ }^{\circ}\text{C}) = 56.7 \text{ g}$

4. What is the specific heat capacity of a substance if 750 J caused 100 grams of it to go from 90 °C to 135 °C?

Answer: $750 \text{ J} = (100 \text{ g})(x)(45 \text{ }^{\circ}\text{C}) = 0.167 \text{ J/g}^{\circ}\text{C}$

5. What would the final temperature be if 500 J are applied to 150 grams of ice at -90 °C?

Answer: $500 \text{ J} = (150 \text{ g})(2.06)(x - (-90 \text{ }^{\circ}\text{C})) = -88.4 \text{ }^{\circ}\text{C}$

6. What would the temperature change by if a 90 gram piece of hot iron cooled by losing 200 J.

Answer: $-200 \text{ J} = (90 \text{ g})(0.449)(x) = -4.95 \text{ }^{\circ}\text{C}$

7. What was the initial temperature if 500 J were applied to 250 g of mercury and the final temperature was 50 °C?

Answer: $500 \text{ J} = (250 \text{ g})(0.140)(50 \text{ }^{\circ}\text{C} - x) = 35.7 \text{ }^{\circ}\text{C}$

Calorimetry

1. What is the specific heat capacity of water?

Answer: $4.18 \text{ J/g}^{\circ}\text{C}$

2. A piece of ice is placed on the sidewalk on a hot summer day. Describe what happens to energy content of the piece of ice, does it gain or lose energy? What happens to the air around the ice, does it gain or lose energy?

Answer: The ice gains energy from the surrounding air and increases its energy (endothermic). The air loses energy to the ice and decreases its energy (exothermic)

Thermodynamics Test Review Answers

Kenneth Rideout, Jonathan Wolf



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