
Practice Problem Set #1- Chapter 2 (Lec 2)**Problem 2.1**

A ship having a mass of 12000 kg is traveling at 60 km/h. How much work must be done to bring it to a stop?

Solution:

a) Using the first law of Thermo I:

$$\Delta U + \Delta KE + \Delta PE = Q + W$$

Assuming:

Isothermal: $\Delta U = 0$

Adiabatic: $Q = 0$

No elevation change: $\Delta PE = 0$

$$W = \Delta KE = \frac{1}{2} m(u_2^2 - u_1^2) = \frac{1}{2} \times 12000 \text{ kg} \left(0^2 - \left[60 \frac{\text{km}}{\text{h}} \times \frac{1000 \text{ m}}{1 \text{ km}} \times \frac{1 \text{ h}}{3600} \right]^2 \right)$$
$$W = -1.7 \times 10^6 \text{ J} = -1700 \text{ kJ}$$

Problem 2.2

One mole of gas in a closed system undergoes a four-step thermodynamic cycle. Use the data given in the following table to determine numerical values for the missing quantities, i.e. fill in the blanks.

Step	$\Delta U^s \text{ (J)}$	$Q \text{ (J)}$	$W \text{ (J)}$
1→2	-100	?	-7000
2→3	?	-2800	?
3→4	?	-700	200
4→1	3500	?	?
1→2→3→4→1	?	?	-1200

Solution:

a) For this system, using the first law of Thermo I, and assuming $\Delta KE = \Delta PE = 0$

$$\Delta U = Q + W$$

Step 1→2:

$$\Delta U^{1 \rightarrow 2} = Q + W \rightarrow -100 = Q - 7000 \rightarrow Q^{1 \rightarrow 2} = 6900 \text{ J}$$

Step 3→4:

$$\Delta U^{3 \rightarrow 4} = Q + W = -700 + 200 = -500 \text{ J}$$

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