

Annotated periodic table of the elements

1 H1.008Hydrogen		2 He4.003Helium		Group name* 1 Hydrogen & the alkali metals 2 Alkaline earth metals 11 Coinage metals 12 Volatile metals 13 Lanthanides 14 Crystallogens 15 Pnictogens 16 Chalcogens 17 Halogens 18 Helium group *Groups 1 to 10 are named after their first members i.e. Group 1 is the Scandium Group																13 B10.81Boron		14 C12.01Carbon		15 N14.01Nitrogen		16 O16.00Oxygen		17 F18.99Fluorine		18 Ne20.18Neon																									
3 Li6.94Lithium		4 Be9.012Beryllium		Most common oxidation states, electronegativity (Mulliken-Pauling)																5 Al26.98Aluminium		6 Si28.09Silicon		7 P30.97Phosphorus		8 S32.06Sulfur		9 Cl35.45Chlorine		10 Ar39.95Argon																									
11 Na22.99Sodium		12 Mg24.31Magnesium		Atomic number and symbol Atomic weight and name																19 K39.10Potassium		20 Ca40.08Calcium		21 Sc44.96Scandium		22 Ti47.87Titanium		23 V50.94Vanadium		24 Cr52.00Chromium		25 Mn54.94Manganese		26 Fe55.85Iron		27 Co58.93Cobalt		28 Ni58.69Nickel		29 Cu63.55Copper		30 Zn65.38Zinc		31 Ga69.72Gallium		32 Ge72.63Germanium		33 As74.92Arsenic		34 Se78.96Selenium		35 Br79.90Bromine		36 Kr83.80Krypton	
37 Rb85.47Rubidium		38 Sr87.62Strontium		57 to 70																39 Y88.91Yttrium		40 Zr91.22Zirconium		41 Nb92.91Niobium		42 Mo95.94Molybdenum		43 Tc98.91Technetium		44 Ru101.07Ruthenium		45 Rh101.07Rhodium		46 Pd106.32Palladium		47 Ag107.87Silver		48 Cd112.41Cadmium		49 In114.82Indium		50 Sn118.71Tin		51 Sb121.76Antimony		52 Te127.60Tellurium		53 I126.91Iodine		54 Xe131.29Xenon					
55 Cs132.91Cesium		56 Ba137.33Barium		89 to 102																57 La138.91Lanthanum		71 Lu174.97Lutetium		72 Hf178.49Hafnium		73 Ta180.95Tantalum		74 W183.84Tungsten		75 Re186.21Rhenium		76 Os190.23Osmium		77 Ir192.22Iridium		78 Pt195.08Platinum		79 Au196.97Gold		80 Hg200.59Mercury		81 Tl204.38Thallium		82 Pb207.2Lead		83 Bi208.98Bismuth		84 Po209Polonium		85 At210Astatine		86 Rn222Radon			
87 Fr223Francium		88 Ra226Radium		103 La138.91Lanthanum																104 Rf181Rutherfordium		105 Db185Dubnium		106 Sg271Seaborgium		107 Bh270Bohrium		108 Hs277Hassium		109 Mt276Meitnerium		110 Ds285Darmstadtium		111 Rg284Roentgenium		112 Cn285Copernicium		113 Nh286Nihonium		114 Fl289Flerovium		115 Mc288Moscovium		116 Lv293Livermorium		117 Ts294Tennessine		118 Og294Oganesson							

Group names*

- Hydrogen & the alkali metals
- Alkaline earth metals
- Coinage metals
- Volatile metals
- Halogens
- Chalcogens
- Posthalogens
- Chalcogens
- Halogens
- Helium group

*Groups 3 to 10 are named after their first members i.e. Group 3 is the Scandium group

Main group
oxidation states
electronegativity
Mendeleev

Atomic number
and symbol
Atomic weight
and name

CLASS COUNTERPARTS

Metals	Nonmetals
Pre-transition	
Active	Active
Transition	Intermediate
Post-transition	Metals
Metals (Ru-Pt, Os-Au)	Metals

Active metals

- Pre-transition metals**
Light, reactive, mostly of low mechanical strength, melting and boiling points. The group 1 members are easily cut with a knife. Only Be, Mg and Al have any structural uses. Most form basic oxides (Be and Al are amphoteric).
- Rare earths**
Closely resemble Ca, Sr and Ba, but many are heavier and they form mostly pale coloured compounds. Similar to one another, esp. La to Lu (aka the lanthanides), and hard to separate for that reason. Abundant but widely dispersed; commercially viable concentrations are rare. Oxides are strongly to moderately basic. Sc, Yb and Lu exhibit weak amphotericism.
- Actinides**
Soft, dense and reactive metals. Those from Ra onwards are synthetic. Early actinides (Th-Am) show some similarities to transition metals and have basic, amphoteric, or acidic oxides; late actinides are more like lanthanides. Many actinides form coloured compounds.

Transition metals

- Moderate to high density metals, with high melting and boiling points, many have high hardness and mechanical strength. Those from Ir onwards are synthetic. Chemically show variable valency and strong tendency to form coordination and colored compounds. Oxides are basic, amphoteric or acidic, depending on oxidation state.
- Noble metals**
Resistant to corrosion and oxidation, unlike base metals such as Fe, Ni, Cu and Zn. Commonly found together in native form. Good catalytic properties. Very easily electroplated; insoluble oxides.
- Frontier metals**
Soft (or brittle), low strength, and with melting points lower than transition metals. Crystalline structures tend to show directional bonding, with generally greater complexity or lower nearest neighbours than other metals. All used or tried as soft metal lubricants. Most are chemically weak; a minority are moderately electropositive. Characterised, to varying degrees, by covalent bonding tendencies, amphotericism, and anionic species such as zincates, stannates, and bismuthates.

Metals

- Metalloids**
Metallic looking brittle, fair electrical conductors. All known in semiconducting forms. Moderate ionisation energies (750–1,000 kJ/mol), and electronegativity values (1.8–2.3). Mostly behave chemically as non-metals. Can form alloys with metals. Amphoteric or weakly acidic oxides.
- Intermediate nonmetals**
H and N are colorless gases; C, P, and Se are metalloids looking solids while S is yellow. The latter four are brittle but all known in malleable, plastic or ductile forms. Moderate to high ionisation energies, low to high electron affinities, moderate to high electronegativity values, relatively poor to moderately strong oxidising agents, tend to form covalent or metallic to near-metallic compounds with metals. Their oxides are acidic, or neutral (H₂O, CO, NO, N₂O).
- Group 1**
The alkali metals (Li, Na, K, Rb, Cs, Fr) and the alkaline earth metals (Be, Mg, Ca, Sr, Ba, Ra) are the most reactive metals in the periodic table. They are soft, shiny, and have low melting points. They react with water to form hydroxides and hydrogen gas. They are also highly reactive with oxygen and halogens.

Nonmetals

- Covalent nonmetals**
O, F, and Cl are, respectively, colorless, pale yellow, and pale yellow-green gases. Br is a red-brown liquid. I is a metallic looking solid. Chemically tend to have high ionisation energies, electron affinities, and electronegativity values, and be relatively strong oxidising agents. Manifestations include oxygen's major association with the ubiquitous processes of corrosion and combustion, and the intrinsically covalent nature of the non-metallic halogens. All five tend to form predominantly ionic compounds with metals. Their oxides are acidic.
- Noble gases**
Colorless, odorless, non-flammable gases with very low chemical reactivity. The first noble gas compound, of approximate composition XePtF₆, was not prepared until 1962. More than half-a-thousand noble gas compounds were synthesised in the ensuing half-century. Compounds of He, Ne, and Ar, in ambient conditions, are not known.

Radioactive
Oganesson is likely to be either a metalloid or a post-transition metal

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R Bogdan



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1. Understanding the eBook Themes On The Periodic Table Answers Key
 - The Rise of Digital Reading Themes On The Periodic Table Answers Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Themes On The Periodic Table Answers Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Themes On The Periodic Table Answers Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Themes On The Periodic Table Answers Key

- Personalized Recommendations
 - Themes On The Periodic Table Answers Key User Reviews and Ratings
 - Themes On The Periodic Table Answers Key and Bestseller Lists
5. Accessing Themes On The Periodic Table Answers Key Free and Paid eBooks
 - Themes On The Periodic Table Answers Key Public Domain eBooks
 - Themes On The Periodic Table Answers Key eBook Subscription Services
 - Themes On The Periodic Table Answers Key Budget-Friendly Options
 6. Navigating Themes On The Periodic Table Answers Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Themes On The Periodic Table Answers Key Compatibility with Devices
 - Themes On The Periodic Table Answers Key Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Themes On The Periodic Table Answers Key
 - Highlighting and Note-Taking Themes On The Periodic Table Answers Key
 - Interactive Elements Themes On The Periodic Table Answers Key
 8. Staying Engaged with Themes On The Periodic Table Answers Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Themes On The Periodic Table Answers Key
 9. Balancing eBooks and Physical Books Themes On The Periodic Table Answers Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Themes On The Periodic Table Answers Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Themes On The Periodic Table Answers Key
 - Setting Reading Goals Themes On The Periodic Table Answers Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Themes On The Periodic Table Answers Key

- Fact-Checking eBook Content of Themes On The Periodic Table Answers Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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