

Answer Key:

1) a. **True**

b. **False** - In the plum pudding model, the atom was envisioned as a sphere of **positive** charge, in which **negatively** charged electrons were randomly distributed.

c. **False** - Rutherford's bombardment experiments with gold foil suggested that alpha particles were deflected by approaching a **positively** charged atomic nucleus.

d. **True**

e. **False** - Atoms that have the same number of **protons** but different number of **neutrons** are called isotopes.

2) $^{90}_{38}\text{Sr}$

Protons = 38 Neutrons = 90-38 = 52 Electrons = 38*

*Assume based on isotope notation shown without a charge

3)

Name	Symbol	Atomic Number	Mass Number	Neutrons	Electrons
Sodium	$^{23}_{11}\text{Na}$	11	23	12	11
Nitrogen	$^{15}_7\text{N}$	7	15	8	7
Barium	$^{136}_{56}\text{Ba}$	56	136	80	56
Lithium	^7_3Li	3	7	4	3
Boron	$^{11}_5\text{B}$	5	11	6	5

Theory Of Atomic Structure Answer Key

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