

Exercises

14.1-14.3 Temperature, Internal energy and states of matter

- Molecules (of the liquid) are in random / haphazard motion (AM)
Molecules (of liquid) are smaller than pollen grains.
 - increase the temperature (of the liquid)
- The sum of (the random distribution of) the KE and PE of (the) molecules.
- No change in KE
Because temperature is constant (during melting)
PE of (the molecules) increases (during melting)
The internal energy increases.

14.4 Specific Heat Capacity

- 35 700 J
- 21 000 J/g°C
- 0.375 kg
- 60 s
- Electrical energy supplied = Heat taken in by water
 $2.5^{\circ}\text{C} = 20^{\circ}\text{C} - 18^{\circ}\text{C}$
 $2100 \times 2 = 5.5 \times 4200 \times 20$
 $E = 92400 \text{ J}$
 - The kinetic energy of the particle is increasing. [1]
- $2 = 3 \times 60 = 180 \text{ s}$

Thermal Energy And Thermodynamics Test Review Answers

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