

Thermochem and Collision theory Study Guide

What is enthalpy? (define it and indicate its symbol)

heat content of a reaction or process. H

What is thermochemistry?

Study of heat (E) changes of a reaction

What is an exothermic reaction?

a reaction that gives off heat

Explain why an exothermic reaction have a negative enthalpy change?

b/c reactants lose heat to become products. (The overall ΔH is negative)

What is an endothermic reaction?

a reaction that absorbs heat

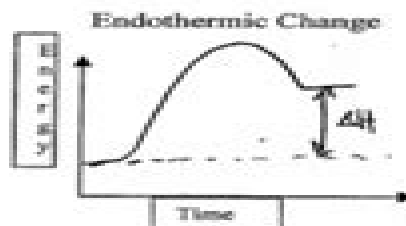
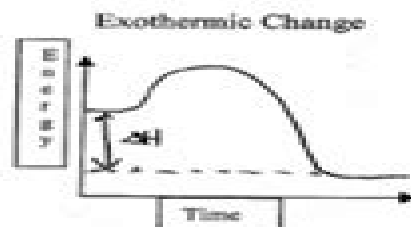
Explain why an endothermic reaction have a positive enthalpy change?

b/c reactants absorb heat to become products (The overall ΔH is +)

Complete the following chart:

Type of Reaction	Sign of ΔH_{rxn}	Which has more energy: reactants or products?	What temperature change would you see?
Exothermic	-	Reactants	Temp would $\uparrow$
Endothermic	+	Products	Temp. would $\downarrow$

Make enthalpy diagrams for the two chemical reactions from above, showing clearly the amount of energy released or gained.



Does the energy go from the surrounding to the chemicals or from the chemicals to the surroundings in an exothermic reaction? from chemicals to surroundings

Does the energy go from the surrounding to the chemicals or from the chemicals to the surroundings in an endothermic reaction? from surroundings to chemicals

Classify each of the following as an exothermic or endothermic process.

Melting ice cubes endothermic

Burning a candle exothermic

Evaporation of water endothermic

Baking Bread endothermic

Splitting a gas molecule apart endothermic (breaking bond!)

Formation of snow in clouds exothermic
(freezing!)

What type of reactions are the following based on the enthalpy?

$C + H_2O \rightarrow CO + H_2$ $\Delta H_{rxn} = +113 \text{ kJ}$ endothermic

$2H_2O_2 \rightarrow 2H_2O + O_2$ $\Delta H = -190 \text{ kJ}$ exothermic

$6CO_2 + 6H_2O + 2803 \text{ kJ} \rightarrow C_6H_{12}O_6 + 6O_2$ endothermic

$3NO_2 + H_2O \rightarrow 2HNO_3 + NO + 138 \text{ kJ}$ exothermic

$P_4 + 5O_2 \rightarrow 3013 \text{ kJ} + P_4O_{10}$ exothermic

Thermo Study Guide Key

Judd E. Hollander

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 - The Rise of Digital Reading Thermo Study Guide Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Thermo Study Guide Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Thermo Study Guide Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Thermo Study Guide Key
 - Personalized Recommendations
 - Thermo Study Guide Key User Reviews and Ratings
 - Thermo Study Guide Key and Bestseller Lists
5. Accessing Thermo Study Guide Key Free and Paid eBooks
 - Thermo Study Guide Key Public Domain eBooks
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 - Thermo Study Guide Key Budget-Friendly Options

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 - Highlighting and Note-Taking Thermo Study Guide Key
 - Interactive Elements Thermo Study Guide Key
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 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Thermo Study Guide Key
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 - Creating a Diverse Reading Collection Thermo Study Guide Key
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 - Minimizing Distractions
 - Managing Screen Time
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 - Setting Reading Goals Thermo Study Guide Key
 - Carving Out Dedicated Reading Time
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 - Fact-Checking eBook Content of Thermo Study Guide Key
 - Distinguishing Credible Sources
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 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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- Interactive and Gamified eBooks

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