

THERMOCHEMISTRY CALCULATIONS WORKSHEET 1

Using reaction equation ratios

- How much heat will be released when 6.44 g of Sulfur reacts with excess O_2 according to the following equation?
 $2S + 3O_2 \rightarrow 2SO_3 \quad \Delta H^\circ = -791.4 \text{ kJ}$
- How much heat will be released when 4.72 g of Carbon reacts with excess O_2 according to the following equation?
 $C + O_2 \rightarrow CO_2 \quad \Delta H^\circ = -393.5 \text{ kJ}$
- How much heat will be absorbed when 38.2 g of Bromine reacts with excess H_2 according to the following equation?
 $H_2 + Br_2 \rightarrow 2HBr \quad \Delta H^\circ = 72.80 \text{ kJ}$
- How much heat will be released when 1.48 g of Chlorine reacts with excess phosphorus according to the following equation?
 $2P + 5Cl_2 \rightarrow 2PCl_5 \quad \Delta H^\circ = -886 \text{ kJ}$
- How much heat will be released when 4.77 g of ethanol (C_2H_5OH) reacts with excess O_2 according to the following equation?
 $C_2H_5OH + 3O_2 \rightarrow 2CO_2 + 3H_2O \quad \Delta H^\circ = -1366.7 \text{ kJ}$
- How much heat will be absorbed when 13.7 g of Nitrogen reacts with excess O_2 according to the following equation?
 $N_2 + O_2 \rightarrow 2NO \quad \Delta H^\circ = -180 \text{ kJ}$
- How much heat will be released when 11.8 g of Iron reacts with excess O_2 according to the following equation?
 $3Fe + 2O_2 \rightarrow Fe_3O_4 \quad \Delta H^\circ = -1120.5 \text{ kJ}$
- How much heat will be released when 18.6 g of Hydrogen reacts with excess O_2 according to the following equation?
 $2H_2 + O_2 \rightarrow H_2O \quad \Delta H^\circ = -571.6 \text{ kJ}$
- How much heat will be transferred when 14.9 g of ammonia reacts with excess O_2 according to the following equation?
 $4NH_3 + 5O_2 \rightarrow 4NO + 6H_2O \quad \Delta H^\circ = -1170 \text{ kJ}$
- How much heat will be transferred when 5.81 g of graphite reacts with excess H_2 according to the following equation?
 $6C(\text{graphite}) + 3H_2 \rightarrow C_6H_6 \quad \Delta H^\circ = 49.03 \text{ kJ}$

Using $\Delta H = mC\Delta T$

- How many kilojoules of heat energy are required to heat all the aluminum (C_p of Al = $.902 \text{ J/g}^\circ\text{C}$) in a roll of aluminum foil (500.0 g) from room temperature (25.0°C) to the temperature of a hot oven (250.0°C)?
- One way to cool down your cup of coffee is to plunge an ice-cold piece of aluminum into it. Suppose you store a 20.0 g piece of aluminum (C_p of Al = $.902 \text{ J/g}^\circ\text{C}$) in the refrigerator at 4.40°C and then drop it into your coffee. The coffee temperature drops from 90.0°C to 55.0°C . How many joules of heat energy did the aluminum block absorb? (Ignore the cooling of the cup)
- Suppose you pick up a 16.0-kilogram ball of iron (such as a "shot-put" ball at a track event). The iron ($C_p = .451 \text{ J/g}^\circ\text{C}$) ball has the same temperature as the atmosphere on a cool day say 16.0°C . How many kilojoules of heat energy must the iron ball absorb to reach the temperature of your body (37.0°C)?
- The specific heat of silver is $0.24 \text{ J/g}^\circ\text{C}$. How much heat in joules must be added to a silver block of mass 86.0 g to raise its temperature by 9.00°C ?
- An 18.7 g sample of platinum metal increases in temperature by 2.30°C when 5.70 J of heat are added. What is the specific heat of platinum?

Answers: 1) -79.4kJ 2) -155kJ 3) 17.4kJ 4) -3.75kJ 5) -142kJ 6) -88.1kJ 7) -78.7kJ 8) -2660kJ 9) -256kJ 10) 3.96kJ
 11) 101kJ 12) 91.2J 13) 152kJ 14) 186J 15) $.132 \text{ J/g}^\circ\text{C}$

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Cady, Carmen Giunta, 2014 The Student Solutions Manual to accompany Atkins Physical Chemistry 10th edition provides full worked solutions to the exercises and the odd numbered discussion questions and problems presented in the parent book. The manual is intended for students and instructors alike and provides helpful comments and friendly advice to aid understanding.

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Preliminary Report on the Thermodynamic Properties of Selected Light Element and Some Related Compounds United States. National Bureau of Standards, 1969

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Chemistry-vol-I Dr Vijay Sarda, Dr A C Handa, Dr K K Arora, A text book on Chemistry

Students Solutions Manual to Accompany Physical Chemistry: Quanta, Matter, and Change 2e Charles Trapp, Marshall Cady, Carmen Giunta, 2014 The Students Solutions Manual to Accompany Physical Chemistry Quanta Matter and Change 2e provides full worked solutions to the a exercises and the odd numbered discussion questions and problems presented in the parent book The manual is intended for students and instructors alike and provides helpful comments and friendly advice to aid understanding

Crystallizing Ideas - The Role of Chemistry Ponnadurai Ramasami, Minu Gupta Bhowon, Sabina Jhaumeer Laulloo, Henri Li Kam

Wah,2016-06-29 Twenty three carefully selected peer reviewed contributions from the International Conference on Pure and Applied Chemistry ICPAC 2014 are featured in this edited book of proceedings ICPAC 2014 a biennial meeting was held in Mauritius in June 2014 The theme of the conference was Crystallizing Ideas The Role of Chemistry and it matched the declaration of the year 2014 as the International Year of Crystallography ICPAC 2014 was attended by 150 participants from 30 countries The chapters in this book reflect a wide range of fundamental and applied research in chemistry and interdisciplinary subjects Crystallizing Ideas The Role of Chemistry is written for graduates postgraduates researchers in industry and academia who have an interest in the fields ranging from fundamental to applied chemistry *A Text Book of Thermo-chemistry and Thermodynamics* Otto Sackur,1917 IIT Chemistry-II ,

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