

Answers - Exercises (Thermochemistry)

Q1: Calculate the specific heat capacity of copper given that 204.75 J of energy raises the temperature of 15g of copper from 25°C to 60°C. (Ans : $c = 0.39 \text{ J/g.K}$)

Answer:

$$q = mc\Delta T$$

$$q = 204.75 \text{ J}$$

$$m = 15\text{g}$$

$$\Delta T = (60 - 25) = 35^\circ\text{C}$$

$$c = \frac{q}{m\Delta T}$$

$$= \frac{204.75 \text{ J}}{(15\text{g})(35^\circ\text{C})}$$

$$= 0.39 \text{ J g}^{-1} \text{ }^\circ\text{C}^{-1}$$

Q2: A chemist burns 0.8650 g of graphite (a form of carbon) in a new bomb calorimeter, and CO_2 forms. If 393.5 kJ of heat is released per mole of graphite and T increases 2.613 K, what is the capacity of the bomb calorimeter? (Ans: 10.84 kJ/K)

Answer:



$$\text{Mass of Carbon (Graphite)} = 0.8650 \text{ g}$$

$$\text{Atomic mass of Carbon} = 12 \text{ g/mol}$$

$$\text{Mol of graphite} = \text{mass/atomic mass} = 0.8650 \text{ g/} 12\text{g mol}^{-1} = 0.072 \text{ mol}$$

$$\text{Heat of combustion, } \Delta H_{\text{comb}} = -393.5 \text{ kJ/mol}$$

$$\text{Temperature change, } \Delta T = 2.613 \text{ K}$$

$$\text{Calorimeter heat capacity, } C_{\text{cal}} = ?$$

$$\Delta H_{\text{comb}} = \frac{q_{\text{rxn}}}{n} = \frac{C_{\text{cal}} \times \Delta T}{n}$$

$$393.5 = \frac{C_{\text{cal}}(2.613)}{0.072}$$

$$C_{\text{cal}} = 10.84 \text{ kJ/K}$$

Thermochemistry Exercises Answers

TD Snyder



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