

Practice Questions: Thermochemistry

- Calculate the work energy, w , gained or lost by the system when a gas expands from 15 L to 35 L against a constant pressure of 1.5 atm.
(1 L·atm = 100 J)
- What is expected when the reaction shown below takes place in a thermally-insulated container outfitted with a movable piston at a constant atmospheric pressure of 1 atm?

$$2\text{C}_2\text{H}_2(g) + 7\text{O}_2(g) \rightarrow 4\text{CO}_2(g) + 6\text{H}_2\text{O}(g)$$
 - Volume will decrease and work will be done on the system.
 - Volume will decrease and work will be done by the system.
 - Volume will increase and work will be done by the system.
 - Volume will decrease and work will be done on the system.
- An ideal gas expands into a vacuum (external pressure = 0) without gaining or losing heat. For this expansion...
 - ΔE increases.
 - ΔE decreases.
 - ΔE does not change.
- Which of the following is TRUE if $E_{\text{sys}} = -95 \text{ J}$?
 - The system is gaining 95 J, while the surroundings are losing 95 J.
 - The system is losing 95 J, while the surroundings are gaining 95 J.
 - Both the system and the surroundings are gaining 95 J.
 - Both the system and the surroundings are losing 95 J.
 - None of the above is true.
- Which of the following signs for q and w represent a system that is doing work on the surroundings, as well as losing heat to the surroundings?
 - $q = -$, $w = -$
 - $q = -$, $w = +$
 - $q = +$, $w = +$
 - $q = +$, $w = -$
 - None of these represent the system referenced above.
- In which case is the work on the system always positive?

A. $\Delta V < 0$	D. $\Delta V = 0$
B. $\Delta E > 0$	E. $\Delta E = 0$
C. $\Delta V > 0$	F. $\Delta E < 0$
- What is the change in internal energy for a system that absorbs 15.0 kJ of heat and performs 610 J of work on the surroundings?
- Calculate the change in internal energy (ΔE) for a system that is giving off 25.0 kJ of heat and is changing from 12.00 L to 6.00 L in volume at 1.50 atm pressure.
(Remember that 101.3 J = 1 L·atm.)
- Find ΔE° for the reaction below if the process is carried out at a constant pressure of 1.00 atm and ΔV (the volume change) equals -24.5 L .
(1 L·atm = 101 J)

$$2\text{CO}(g) + \text{O}_2(g) \rightarrow 2\text{CO}_2(g) \quad \Delta H^\circ = -566 \text{ kJ}$$
- Which of the following results in the largest decrease (most negative change) in internal energy for a system?
 - The system volume decreases and the surroundings get colder.
 - The system volume decreases and the surroundings get hotter.
 - The system volume increases and the surroundings get colder.
 - The system volume increases and the surroundings get hotter.
 - The system volume remains constant and the surroundings get colder.

Thermochemistry Practice Multiple Choice Questions

Julia Schneider



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