

Thermochemistry Practice Problems ANSWER KEY

1. Calculate the standard enthalpy of reaction for the process



$$\Delta H = -176 \text{ kJ}$$

2. Calculate the heat evolved in the combustion of ethyl alcohol, ΔH_{comb}



$$\Delta H = -1366.7 \text{ kJ}$$

3. Calculate the heat required to decompose limestone, CaCO_3 , to lime, CaO , and carbon dioxide, with all substances at standard conditions.



$$\Delta H^\circ_f \text{ CaCO}_3(\text{s}) = -1206.9 \text{ kJ/mol}$$

$$\Delta H^\circ_f \text{ CaO}(\text{s}) = -635.1 \text{ kJ/mol}$$

$$\Delta H^\circ_f \text{ CO}_2(\text{g}) = -393.5 \text{ kJ/mol}$$

$$\Delta H = 178.3 \text{ kJ}$$

4. $\text{CH}_4(\text{g}) + 3 \text{Cl}_2(\text{g}) \rightarrow \text{CHCl}_3(\text{g}) + 3 \text{HCl}(\text{g}) \quad \Delta H = ?$



$$\Delta H = -305.2 \text{ kJ}$$

5. Cyanamide, CH_2N_2 , is a weak acid that is sometimes used as a fertilizer. Calculate the standard enthalpy of formation for cyanamide, given the following standard enthalpies of reaction:



$$\Delta H = 62.1 \text{ kJ}$$

6. The standard molar enthalpy change is -905.2 kJ for the oxidation of ammonia.



Calculate the standard molar enthalpy of formation for ammonia based on the following standard enthalpies of formation: $\Delta H^\circ_f \text{ NO}(\text{g}) = 90.3 \text{ kJ/mol}$ and $\Delta H^\circ_f \text{ H}_2\text{O}(\text{g}) = -241.8 \text{ kJ/mol}$

$$-46.1 \text{ kJ/mol}$$

7. The thermite reaction: $2 \text{Al}(\text{s}) + \text{Fe}_2\text{O}_3(\text{s}) \rightarrow \text{Al}_2\text{O}_3(\text{s}) + 2 \text{Fe}(\text{s}) + 852 \text{ kJ}$

If you begin with 10.0 g of Al and excess Fe_2O_3 , how many kJ of heat are evolved at constant pressure?

$$158 \text{ kJ}$$

Thermochemistry Practice Problem Answers

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Chemistry: 1001 Practice Problems For Dummies (+ Free Online Practice) Heather Hattori, Richard H.

Langley, 2022-05-10 Practice your way to a better grade in your Chemistry class Chemistry 1001 Practice Problems For Dummies gives you 1 001 opportunities to practice solving problems on all the topics covered in your chemistry class in the book and online Get extra practice with tricky subjects solidify what you've already learned and get in depth walk throughs for every problem with this useful book These practice problems and detailed answer explanations will catalyze the reactions in your brain no matter what your skill level Thanks to Dummies you have a resource to help you put key concepts into practice Work through multiple choice practice problems on all Chemistry topics covered in class Step through detailed solutions to build your understanding Access practice questions online to study anywhere any time Improve your grade and up your study game with practice practice practice The material presented in Chemistry 1001 Practice Problems For Dummies is an excellent resource for students as well as parents and tutors looking to help supplement classroom instruction Chemistry 1001 Practice Problems For Dummies 9781119883531 was previously published as 1 001 Chemistry Practice Problems For Dummies 9781118549322 While this version features a new Dummies cover and design the content is the same as the prior release and should not be considered a new or updated product **Survival Guide to General Chemistry**

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Elementary Chemical Reactor Analysis Rutherford Aris, 2013-09-03 Elementary Chemical Reactor Analysis focuses on the processes reactions methodologies and approaches involved in chemical reactor analysis including stoichiometry

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