



Unit: Thermodynamics "Thermodynamic Equations" - Week 8 10/26/2017



Answer these 4 questions using the following problems using dimensional analysis. Show all work.

1. The combustion enthalpy, ΔH_c , of acetylene is known. Use this information to determine combustion enthalpy of acetylene given that when 2.5 moles of acetylene is combusted:



Answer: 12990 kJ



2. Combustion heat for ethanol from the reaction of oxygen, carbon dioxide, and water. How many moles of water can be made from 10.0 moles of ethanol? $\Delta H_c = -1366 \text{ kJ}$



Answer: 15.0 moles

3. Methanol combustion, ΔH_c . Oxygen from the atmosphere, found in ethanol. Ethanol combustion equation for combustion for products is 2 moles of carbon dioxide.



Answer: 170.4 kJ

4. Burning acetylene, C_2H_2 , combustion product is 2 moles. Ethanol combustion with 2 moles of water. Combustion of 10.0 moles of acetylene.



Answer: 12990 kJ

Thermochemistry Review Answer Key Solutions

M Woodhall



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