

Thermochemistry Review Part 1

Definition Questions: you should be able to use or apply the following terms. Be able to write complete definitions for,

energy	temperature	standard state
potential energy	heat	standard heat of formation
kinetic energy	specific heat capacity	exothermic
enthalpy	endothermic	Law of conservation of Energy
thermal	Hess's Law	isolated, closed, open system

1. The specific heat capacity of diamond is $0.5050 \text{ J/g}^\circ\text{C}$. How much energy is required to heat 25.0 g of diamond from 10.5°C to 15.6°C ?

$$q = m \times c \times \Delta t$$

$$= 25 \text{ g} \times 0.5050 \text{ J/g}^\circ\text{C} \times (15.6 - 10.5^\circ\text{C}) = 64.39 \text{ J}$$

2. A piece of aluminum with a mass of 50 g and an initial temperature of 90°C is placed into 100 mL of water at a temperature of 25°C . The temperature of water rises to 31.3°C . Determine the specific heat capacity of aluminum.

$Q \text{ lost (Al)} = Q \text{ gained (Water)}$

$$-(mc\Delta t) = +mc \Delta t$$

$$-(50 \times c \times (31.3 - 90)) = (100 \times 4.18 \times (31.3 - 25))$$

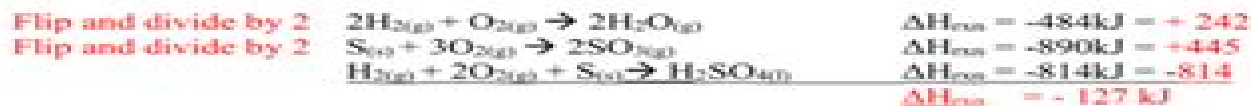
$$-(-2935 c) = 2633.4$$

$$c = 2633.4 / 2935 = 0.897 \text{ J/g}^\circ\text{C}$$

3. How much energy is absorbed by 300 g of methanol, CH_3OH , as it evaporates? Given: The molar heat of vaporization is 35.3 kJ/mol .

$q = ?$	$\Delta H = q/n$	$n \text{ CH}_3\text{OH} = m/\text{MM}$
$m = 300 \text{ g}$	$q = H \times n$	$n = 300 \text{ g} / 32 \text{ g/mol}$
$\Delta H = 35.3 \text{ kJ/mol}$	$q = (35.3 \text{ kJ} \times 9.375 \text{ mol})$	$n = 9.375 \text{ mol}$
	$q = 332.81 \text{ kJ}$	

4. What is the molar enthalpy of the formation of $1 \text{ mol H}_2\text{SO}_{4(l)}$ given the following information?



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Table of Contents Thermochemistry Review Answer Key

1. Understanding the eBook Thermochemistry Review Answer Key
 - The Rise of Digital Reading Thermochemistry Review Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Thermochemistry Review Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Thermochemistry Review Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Thermochemistry Review Answer Key
 - Personalized Recommendations
 - Thermochemistry Review Answer Key User Reviews and Ratings

- Thermochemistry Review Answer Key and Bestseller Lists
- 5. Accessing Thermochemistry Review Answer Key Free and Paid eBooks
 - Thermochemistry Review Answer Key Public Domain eBooks
 - Thermochemistry Review Answer Key eBook Subscription Services
 - Thermochemistry Review Answer Key Budget-Friendly Options
- 6. Navigating Thermochemistry Review Answer Key eBook Formats
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 - Thermochemistry Review Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Thermochemistry Review Answer Key
 - Highlighting and Note-Taking Thermochemistry Review Answer Key
 - Interactive Elements Thermochemistry Review Answer Key
- 8. Staying Engaged with Thermochemistry Review Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Thermochemistry Review Answer Key
- 9. Balancing eBooks and Physical Books Thermochemistry Review Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Thermochemistry Review Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Thermochemistry Review Answer Key
 - Setting Reading Goals Thermochemistry Review Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Thermochemistry Review Answer Key
 - Fact-Checking eBook Content of Thermochemistry Review Answer Key
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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