# Thermochemistry Unit Review Packet

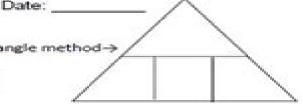
Part 1 Specific Heat Equation Problems

Set up and use the triangle method→

The specific heat equation is  $Q = m \cdot c \cdot \Delta T$ ,  $Q = heat energy m = mass and <math>\Delta T = char$ 

Q = heat energy, m = mass, and  $\Delta T$  = change in temperature ( $T_{finel} - T_{initiel}$ )

Show your work and label all of the units of measurement throughout.



1. How many joules of heat are needed to raise the temperature of 52.0 g of zirconium from 25°C to 56°C, if the specific heat of zirconium is  $0.27 \text{ J/g}^{\circ}$ -C?

2. A 3.65 g piece of iron absorbs 144.7 joules of heat energy and its temperature changes from 23.0°C to 111°C. Calculate the specific heat capacity of iron.

#### Part 2 Gibb's Free Energy Equation Problems

 For the reaction NH<sub>4</sub>Br(s) → NH<sub>3</sub>(g) + HBr(g), at 302.7°K, the enthalpy change is 158 kJ/mol, and the change in entropy is 0.422 kJ/mol·°K. Calculate ΔG.

2. Calculate the entropy of the vaporization phase change for silver at  $2453^{\circ}$ K. The change in enthalpy is +250.58 kJ/mol.

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United States. National Bureau of Standards

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