## Thermochemistry Test Preview

#### Matching

a. calorimeter b. calorie c. joule c. joule f. heat capacity f. capacity of heat needed to change the temperature of I g of a substance by I°C f. quantity of heat needed to change the temperature of an object by I°C f. device used to measure the heat absorbed or released during a chemical or physical process f. heat of needs into with the cornect statement below. heat of reaction heat of formation heat of solution heat of formation heat of formation heat of heat summation  f. the enthalpy change caused by dissolving a substance heat of solution heat of heat summation  f. the enthalpy change caused by dissolving a substance heat of heat summation  f. the enthalpy change is not make accompanies the formation of a compound from its elements heat of needs in enthalpy that accompanies the formation of a compound from its elements heat of reaction to give the final heat of reaction  Multiple Choice  Multiple Choice  Multiple Choice  Multiple Choice  Multiple Choice  A piece of metal is heat in the exhaust.  The energy is fost as heat in the exhaust.  The energy heats the parts of the engine  a. The energy heats the parts of the engine  a. The energy heats the parts of the engine  a. The energy heats the parts of the engine  a. The temperature of the water will increase.  The temperature of the water will increase and the temperature of the metal will increase.  The temperature of the water will increase and the temperature of the metal will decrease.  The temperature of the water will increase and the temperature of the metal will increase.  The temperature of the water will increase and the te			Match each item with the correct statement below	W.,		
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<ul> <li>a. The flask would probably feel cooler than before the reaction started.</li> <li>b. The flask would probably feel warmer than before the reaction started.</li> <li>c. The flask would feel the same as before the reaction started.</li> </ul>		1.5			그리고 있는데 그렇게 되었다면 하는데 그 그리고 있다면 하는데	
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<ul> <li>The flask would feel the same as before the reaction started.</li> </ul>						

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