

Thermochemistry Worksheet Answers

Useful Values:

Heat of Fusion for Water: $\Delta H_{\text{fus}} = 6 \text{ kJ/mol}$ **Common error: not converting to moles**
 Heat of Vaporization for Water: $\Delta H_{\text{vap}} = 45 \text{ kJ/mol}$ **Common error: same**
 1 calorie = 4.18 Joules

Material	Specific Heat in Joules/g/K	in Cal/g/K	
Water (ice)	2.05	0.490	Common error here: using moles instead of grams.
Water (liquid)	4.18	1.00	
Water (gas)	2.08	0.498	

- How much energy does it take to warm the following substances the listed amount?
 - 20g of ice from -45°C to -15°C ? **294 calories / 1230 Joules**
 - 20g of water from 25°C to 55°C ? **600 calories / 2508 Joules**
 - 20g of water vapor from 130°C to 160°C ? **299 calories / 1249 Joules**
- To melt ice, one must bring ice to the freezing point (0°C), and then thaw it into water. How much heat would it take to melt...
 - 30g of ice at 0°C ? **10 kJ (10,000 J)**
 - 30g of ice at -40°C ? **12.46 kJ (12,460 J)**
- To boil water, one must raise it to the boiling point (100°C), and then boil it into steam. How much heat would it take to boil...
 - 40g of water at 100°C ? **100 kJ (100,000 Joules)**
 - 40g of water at 90°C ? **101.7 kJ (101,672 Joules)**
 - 40g of water at 0°C ? **116.7 kJ (116,720 Joules)**
 - 40g of ice at 0°C ? **130.1 kJ (130,053 Joules)**
 - 40g of ice at -80°C ? **136.6 kJ (136,613 Joules)**
- If 700 calories of heat raise 400g of Phonium by 20°C , what is the specific heat of Phonium? **0.0875 cal/g/K**
- How much heat energy is stored in 2g of ice at -30°C ? (remember: use Kelvin) **238 cal / 996 J**
- How much heat energy is stored in 2g of Phonium at 70°C ? (Assume Phonium is a solid at 70°C) **1672 cal**
- If 100g of water at 10°C comes in contact with _____, what will the final temperature be? (Remember, you have the specific heat of Phonium from question 4.)
 - 100g of water at 90°C **50° C**
 - 100g of Phonium at 90°C (Hint: the water and Phonium end at the same temperature) **16.44° C**
Common error: 6.44 degrees is the temperature change, and the water gets warmer.
 - 20g of Phonium at 90°C **11.38° C**
 - 1g of Phonium at 90°C **10.07° C**
- If 450g of Phakelite at 100°C is added to 1000g of water at 20°C , the temperature of the water increases by 5° . What is the specific heat of Phakelite? **0.148 cal/g/K**
- If 450g of Nonsensium at 500°C is added to 1000g of water at 20°C , the temperature of the water increases by 25° . What is the specific heat of Nonsensium? **0.122 cal/g/K**

Thermochemistry Worksheet Answers Unit 1

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Thermochemistry Worksheet Answers Unit 1 Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Thermochemistry Worksheet Answers Unit 1**," written by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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