

■ Review | Constants | Periodic Table

To determine whether a shiny gold-colored rock is actually gold, a chemistry student decides to measure its heat capacity. She first weighs the rock and finds it has a mass of 4.4 g. She then finds that upon absorption of 55.6 J of heat, the temperature of the rock rises from 25 °C to 54 °C.

Substance	Specific heat capacity, C_s (J/(g·°C))*
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Elements

Lead	0.128
Gold	0.129
Silver	0.235
Copper	0.385
Iron	0.449
Aluminum	0.903

Compounds

Ethanol	2.42
Water	4.18

Materials

Glass (Pyrex)	0.75
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You may want to reference (Pages 260 - 264) Section 6.4 while completing this problem.

Part A

Find the specific heat capacity of the substance composing the rock.

Express the specific heat in joules per gram-Celsius to two significant figures.

$$C_s = \text{[input box]} \text{ J/g} \cdot ^\circ\text{C}$$

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Part B

Determine whether the value is consistent with the rock being pure gold by referring to the table.

- ☐ Yes, the value is consistent with the rock being pure gold.
- ☐ No, the value is inconsistent with the rock being pure gold.

Thermochemistry Assessment Answers

Mary L. Kotz



Thermochemistry Assessment Answers:

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