

Unit 4 Thermochemistry Worksheet

- The enthalpy of formation for silver nitrate is -2304 kJ/mol .
 - Write the balanced reaction representing this enthalpy.
 - What is the value of the enthalpy to form three (3) moles of silver nitrate?
 - What is the value of the enthalpy to form 15.490 grams of silver nitrate?

Answer:

- $\text{Ag}(s) + \frac{1}{2} \text{N}_2(g) + \frac{3}{2} \text{O}_2(g) \rightarrow \underline{1} \text{ AgNO}_3(s); \Delta H = -2304 \text{ kJ/mol}$
- $(3 \text{ mol})(-2304 \text{ kJ/mol}) = -6912 \text{ kJ}$
- $\underline{15.490 \text{ g}} \left(\frac{1 \text{ mol}}{169.88 \text{ g}} \right) \left(\frac{-2304 \text{ kJ}}{1 \text{ mol}} \right) = -210.08 \text{ kJ}$

- When 1.00 L of 1.00 M $\text{Ba}(\text{NO}_3)_2$ solution at 25.0°C is mixed with 1.00 L of 1.00 M Na_2SO_4 solution at 25.0°C in a calorimeter, the white solid BaSO_4 forms and the temperature of the mixture increases to 28.1°C . Assuming that the calorimeter absorbs only a negligible quantity of heat, that the specific heat capacity of the solution is $4.18 \text{ J/}^\circ\text{C}\cdot\text{g}$, and that the density of the final mixture is 1.0 g/mL , calculate the enthalpy change per mole of NaNO_3 formed.

Answer:

$$\underline{1} \text{ Ba}(\text{NO}_3)_2 + \underline{1} \text{ Na}_2\text{SO}_4 \rightarrow \underline{2} \text{ NaNO}_3 + \underline{1} \text{ BaSO}_4$$

$$Q_{\text{mixture}} = m_{\text{mixture}} C_{\text{mixture}} \Delta T_{\text{mixture}} = (1000 + 1000 \text{ g})(4.18)(28.1 - 25.0^\circ\text{C}) = +25916 \text{ J}$$

Select one of the reactants to obtain moles:

$$(1.00 \text{ L})(1.00 \text{ M}) = 1.00 \text{ mol Ba}(\text{NO}_3)_2$$

$$\left(\frac{+25916 \text{ J}}{1 \text{ mol Ba}(\text{NO}_3)_2} \right) \left(\frac{1 \text{ mol Ba}(\text{NO}_3)_2}{2 \text{ mol Na}(\text{NO}_3)} \right) = +12958 \text{ J}$$

$$\underline{1.30 \times 10^4 \text{ J}}$$

- The molar heat capacity of mercury is $28.1 \text{ J/mol}\cdot\text{K}$. What is the specific heat capacity of this metal?

Answer:

Specific heat capacity has units of $\text{J/g}\cdot^\circ\text{C}$.

Since the temperature is a **change in temperature**, rather than a direct temperature, the unit is based on the difference between two temperatures. Since $^\circ\text{C}$ and K are converted through addition or subtraction, the difference between two temperatures in $^\circ\text{C}$ is the same difference between the same temperatures in K. Take for instance $T_i = 25^\circ\text{C}$ and $T_f = 55^\circ\text{C}$. The difference between these two is $T_f - T_i = 55 - 25 = 30^\circ\text{C}$. At the same time, these temperatures in Kelvin would appear as $T_i = 298\text{K}$ and $T_f = 328\text{K}$. The difference between these two is $T_f - T_i = 328 - 298 = 30\text{K}$. The magnitude is the same. Therefore, the molar heat capacity of mercury can also be represented as $\text{J/mol}\cdot^\circ\text{C}$.

Now, the moles can be converted to grams:

$$\left(\frac{28.1 \text{ J}}{\text{mol}\cdot^\circ\text{C}} \right) \left(\frac{1 \text{ mol Hg}}{200.59 \text{ g Hg}} \right) = 0.140 \frac{\text{J}}{\text{g}\cdot^\circ\text{C}}$$

Thermochemistry Problems Worksheet Number One

Answeres

Maka Aleksishvili, Shota Sidamonidze

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