

Thermochemistry Worksheet #1

AP Chemistry, Chapter 6

ANSWERS

- 1) +1000 J, endothermic (q is positive)
- 2) -2267 J, work done by system on the surroundings
- 3) $\Delta E = +300$ J, $q = -200$ J exothermic
- 4) 0.5354 kJ, 0.1280 kcal
- 5) $q = +4184$ J; at constant P, so $\Delta H = q$ or $\Delta H = +4184$ J;
 $w = -30.4$ J; $\Delta E = 4154$ J
- 6) $q = +6500$ J, $w = -9620$ J, $\Delta E = -3120$ J
- 7) a. endothermic, 1.8 kJ = 0.43 kcal
b. exothermic, -2.4 kcal = -10 kJ
c. exothermic, -0.025 kcal = -0.10 kJ
d. exothermic, -95.3 kJ = -22.8 kcal
- 8) 0.0329 J/g°C or 0.00786 cal/g°C
- 9) 59.7 kJ
- 10) -49.4 kJ/mol, exothermic
- 11) -289 kJ/mol H_2
- 12) 18.9 L
- 13) -5670 kJ/mol sucrose, -11200 kJ/mol stearic acid
- 14) $\Delta H_{rxn} = -822.3$ kJ
- 15) $\Delta H_{rxn} = -78.7$ kJ

Thermochemistry Questions And Answers

Aftab S. Hassan, James L. Flowers



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Table of Contents Thermochemistry Questions And Answers

1. Understanding the eBook Thermochemistry Questions And Answers
 - The Rise of Digital Reading Thermochemistry Questions And Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Thermochemistry Questions And Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Thermochemistry Questions And Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Thermochemistry Questions And Answers
 - Personalized Recommendations
 - Thermochemistry Questions And Answers User Reviews and Ratings

- Thermochemistry Questions And Answers and Bestseller Lists
- 5. Accessing Thermochemistry Questions And Answers Free and Paid eBooks
 - Thermochemistry Questions And Answers Public Domain eBooks
 - Thermochemistry Questions And Answers eBook Subscription Services
 - Thermochemistry Questions And Answers Budget-Friendly Options
- 6. Navigating Thermochemistry Questions And Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Thermochemistry Questions And Answers Compatibility with Devices
 - Thermochemistry Questions And Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Thermochemistry Questions And Answers
 - Highlighting and Note-Taking Thermochemistry Questions And Answers
 - Interactive Elements Thermochemistry Questions And Answers
- 8. Staying Engaged with Thermochemistry Questions And Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Thermochemistry Questions And Answers
- 9. Balancing eBooks and Physical Books Thermochemistry Questions And Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Thermochemistry Questions And Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Thermochemistry Questions And Answers
 - Setting Reading Goals Thermochemistry Questions And Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Thermochemistry Questions And Answers
 - Fact-Checking eBook Content of Thermochemistry Questions And Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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