

CHAPTER 7

THERMOCHEMISTRY SHORT QUESTION AND ANSWERS

Q.1 Define the following terms and give three examples of each.

Ans.

(i) System:

The substance which is under experiment or under observation is called as system.

Examples:

(i) $\text{Pb}(\text{NO}_3)_2$ in decomposition of $\text{Pb}(\text{NO}_3)_2$.

(ii) Zn and CuSO_4 solution, the reaction mixture in the vessel.

(iii) CaCO_3 in thermal decomposition of CaCO_3

(ii) Surroundings:

Everything around the system which is not a part of system is called surroundings.

For example

During the reaction between Zn and CuSO_4 solution vessel and air etc are surroundings.

(iii) State function:

A macroscopic property of a system which has some definite value for initial and final state and independent of the path followed e.g.

(i) Pressure (ii) Temperature (iii) Internal energy.

Note: Heat is not a state function.

Q.2 Describe the units of energy.

Ans.

Mostly Joule and calorie are used for the measurement of energy.

Calorie:

The amount of energy required to raise the temperature of one gram of water from 14.5°C to 15.5°C is called one calorie.

Joule:

Thermochemistry Short Answers

Peter William Atkins, Julio de Paula



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Thermochemistry Mr. Ajeet Singh, 2018-05-16 Introduces thermodynamic principles related to chemical processes including enthalpy entropy Gibbs free energy and calorimetry

Thermochemistry and Its Applications to Chemical and Biochemical Systems M.A.V. Ribeiro Da Silva, 2012-12-06 Proceedings of the NATO Advanced Study Institute on Thermochemistry Today and Its Role in the Immediate Future Viano do Castelo Portugal July 5 15 1982

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Chimie Pure Et Appliquée, 1964 Some vols include Plenary lectures of the International Conference on Organometallic Chemistry

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