

Answer Key

1

Thermodynamics.

1. $\gamma = 1.5$; $V_1 = 1600 \text{ c.c.}$; $V_2 = 400 \text{ c.c.}$
 $P_1 = 150 \text{ kPa} = 150 \times 10^3 \text{ Pa.}$ $P_2 = ?$

adiabatic Process $P_2 V_2^\gamma = P_1 V_1^\gamma$

$$P_2 = P_1 \left(\frac{V_1}{V_2} \right)^\gamma = 150 \left(\frac{1600}{400} \right)^{1.5}$$
$$= 150 \times (4)^{1.5} = 150 \times 2^{2 \times \frac{3}{2}} = 150 \times 2^3$$
$$= 150 \times 8 = 1200 \text{ kPa.}$$

Work done on the gas in the Process

$$W = \frac{P_1 V_1 - P_2 V_2}{\gamma - 1} = \frac{(150 \times 10^3)(1600 \times 10^{-6}) - (1200 \times 10^3)(400 \times 10^{-6})}{1.5 - 1}$$

$$W = \frac{240 - 480}{0.5} = -480 \text{ J.}$$

2. $T = 127^\circ \text{C} = 127 + 273 = 400 \text{ K}$ $V_2 = 2V_1$

$$W = 2.3026 RT \log_{10} \frac{V_2}{V_1} = 2.302 \times 8.31 \times 400 \log_{10} 2$$
$$W = 2.303 \times 10^3 \text{ J.}$$

Amount of heat absorbed $= \frac{2.302 \times 10^3}{1.5}$
 $= 648.5 \text{ cal.}$

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Neil E. Schore



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