

First law of thermodynamics for various processes:

- i. **Isothermal process:** In such process, temperature is constant, internal energy is constant, hence $\Delta U = 0$

Therefore $\Delta U = Q + W$

$$0 = Q + W$$

Or $W = -Q$

- ii. **Adiabatic process:** In this process, there is no exchange of heat between system and its surroundings

i.e. $Q = 0$, then

$$-\Delta U = -W$$

- iii. **Isochoric process:**

we know,

$$W = -P_{\text{ext}} \Delta V$$

$$\Delta U = Q + W$$

$$\Delta U = Q - P_{\text{ext}} \Delta V$$

In isochoric process

$$\Delta V = 0$$

Therefore $\Delta U = Q_v$

U is state function, Q_v is also a state function.

RIZWANA
MOHAMMED

Thermodynamics 12th Std

Michael Seilmaier



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