Chemistry 104 Final Exam With Completely Detailed Solutions 2024

How many moles of hydroxide ions are there in 25 mL of 0.010M solution of Ba(OH)2?

a. 7.5 x 10-1 mol

b. 5.0 x 10-1 mol

c. 2.5 x 10-1 mol

d. 5.0 x 10-4 mol

e. 2.5 x 10-4 mol - correct answer.d. 5.0 x 10-4 mol

What is the molarity of a solution that contains 46 grams of Na2SO4 (Mw = 142 g/mol) dissolved in enough water to make 750 mL of solution?

a. 2.3M

b. 0.60M

c. 0.52M

d. 0.43M

e. 0.19M - correct answer.d. 0.43M

Imagine that you go into the lab and measure 35 mL of a 0.60M stock solution. You pour the 35 mL of stock solution into a beaker and dilute the solution by adding another 35 mL of water. What is the molarity of the new solution?

a. 1.2M

b. 0.60M

c 0.30M

d. 0.15M

e. 0.040M - correct answer.c. 0.30M

What is the correct expression for the equilibrium constant, K, for the following reaction? P4(s) + 5O2(g) = P4O10(s) - correct answer.K= 1/[O2]^5

If for a certain gas phase reaction mixture Q = 1, what can you say about the reaction?

The reaction will proceed towards the product to achieve equilibrium.

The equilibrium position will be determined by the value of K.

The reaction will proceed to the left to achieve equilibrium.

d. The reaction is already at equilibrium.

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Katrin Zwirglmaier

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