

Model 2 – Ternary Acids (Oxyacids)

Acid	Name of Acid in Aqueous Solution	Cation (+)	Polyatomic Anion (-)	Polyatomic Anion Name
HClO_3	Chloric acid	H_3O^+	ClO_3^{-1}	Chlorate
H_2SO_3	Sulfurous acid	$2\text{H}_3\text{O}^+$	SO_3^{-2}	Sulfite
H_2SO_4	Sulfuric acid	$2\text{H}_3\text{O}^+$	SO_4^{-2}	Sulfate
H_3PO_3	Phosphorous acid	$3\text{H}_3\text{O}^+$	PO_3^{-3}	Phosphite
H_3PO_4	Phosphoric acid	$3\text{H}_3\text{O}^+$	PO_4^{-3}	Phosphate
HNO_3	Nitric acid	H_3O^+	NO_3^{-1}	Nitrate
HNO_2	Nitrous acid	H_3O^+	NO_2^{-1}	Nitrite
H_2CO_3	Carbonic acid	$2\text{H}_3\text{O}^+$	CO_3^{-2}	Carbonate

6. Look at the formulas of the ternary acids in Model 2.

a. How are ternary acids different from binary acids in their structure?

Ternary acids contain three elements and they also contain the element oxygen.

b. What number do you think the prefix "ter-" refers to?

"ter" refers to three

7. When ternary acids are mixed with water, ions will form. Fill in the table above with the formulas and names of the anions.

8. Examine the pairs of ternary acids in Model 2 that contain sulfur, phosphorus, and nitrogen. Each pair has one acid that ends in "-ic" and another that ends in "-ous." These endings are related to the name of the polyatomic anion found in the acid ("-ate" or "-ite"). Complete the statements below with the correct acid name ending.

Polyatomic anion ending is "-ate" → acid name ending is "ic"

Polyatomic anion ending is "-ite" → acid name ending is "ous"

9. If the prefix "hydro-" were used to name a ternary acid, what problem would this create when naming HClO_3 ?

ClO_3^{-1} has the name of Chlorate. "ate" must be replaced with "ic". If "hydro" is used as a prefix, it would be hydrochloric acid. This has the formula HCl , not HClO_3 .

10. Write a rule for naming ternary acids.

- The prefix "hydro" is not used in the name when an acid contains more than two elements and one of the elements is oxygen.

- "ate" is replaced with "ic" and "ite" is replaced with "ous".

11. Predict the formula for chlorous acid.



12. Circle the acid(s) below that would be named beginning with the prefix "hydro-."



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Table of Contents Solubility Answer Key Pogil

1. Understanding the eBook Solubility Answer Key Pogil
 - The Rise of Digital Reading Solubility Answer Key Pogil
 - Advantages of eBooks Over Traditional Books
2. Identifying Solubility Answer Key Pogil
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Solubility Answer Key Pogil
 - User-Friendly Interface
4. Exploring eBook Recommendations from Solubility Answer Key Pogil
 - Personalized Recommendations
 - Solubility Answer Key Pogil User Reviews and Ratings
 - Solubility Answer Key Pogil and Bestseller Lists
5. Accessing Solubility Answer Key Pogil Free and Paid eBooks
 - Solubility Answer Key Pogil Public Domain eBooks
 - Solubility Answer Key Pogil eBook Subscription Services
 - Solubility Answer Key Pogil Budget-Friendly Options
6. Navigating Solubility Answer Key Pogil eBook Formats

- ePub, PDF, MOBI, and More
- Solubility Answer Key Pogil Compatibility with Devices
- Solubility Answer Key Pogil Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Solubility Answer Key Pogil
 - Highlighting and Note-Taking Solubility Answer Key Pogil
 - Interactive Elements Solubility Answer Key Pogil
- 8. Staying Engaged with Solubility Answer Key Pogil
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Solubility Answer Key Pogil
- 9. Balancing eBooks and Physical Books Solubility Answer Key Pogil
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Solubility Answer Key Pogil
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Solubility Answer Key Pogil
 - Setting Reading Goals Solubility Answer Key Pogil
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Solubility Answer Key Pogil
 - Fact-Checking eBook Content of Solubility Answer Key Pogil
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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