

Determination of the Solubility-Product Constant for a Sparingly Soluble Salt

27

A. Preparation of a Calibration Curve

Initial $[\text{CrO}_4^{2-}]$ 0.0024 M

Volume of 0.0024 M K_2CrO_4	Total volume	$[\text{CrO}_4^{2-}]$	Absorbance
1. <u>1.00 mL</u>	<u>100. mL</u>	<u> </u>	<u>0.060</u>
2. <u>5.00 mL</u>	<u>100. mL</u>	<u> </u>	<u>0.299</u>
3. <u>10.00 mL</u>	<u>100. mL</u>	<u> </u>	<u>0.600</u>
4. <u>15.00 mL</u>	<u>100. mL</u>	<u> </u>	<u>0.840</u>

Molar absorption coefficient for $[\text{CrO}_4^{2-}]$ Average molar absorption coefficient Standard deviation (show calculations)

B. Determination of the Solubility-Product Constant

Absorbance	$[\text{CrO}_4^{2-}]$	$[\text{Ag}^+]$	K_{sp} of Ag_2CrO_4
1. <u>0.720</u>	<u> </u>	<u> </u>	<u> </u>
2. <u>0.750</u>	<u> </u>	<u> </u>	<u> </u>
3. <u>0.780</u>	<u> </u>	<u> </u>	<u> </u>

Solubility Product Lab Answers

**Ralph Petrucci, William
Harwood, Geoffrey Herring**



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