

PART ONE THERMODYNAMICS

CHAPTER 1: First Law of Thermodynamics

- ✓ 1.1 Assuming that the atmosphere is isothermal at 0 °C and that the average molar mass of air is 29 g mol⁻¹, calculate the atmospheric pressure at 20,000 ft above sea level.

SOLUTION

$$h = (2.0 \times 10^4 \text{ ft})(12 \text{ in ft}^{-1})(2.54 \text{ cm in}^{-1})(10^{-2} \text{ m cm}^{-1}) \\ = 6096 \text{ m}$$

$$P = P_0 e^{-gMh/RT}$$

$$P = (1.013 \text{ bar}) \exp \left[\frac{-(9.8 \text{ m s}^{-2})(29 \times 10^{-3} \text{ kg mol}^{-1})(6096 \text{ m})}{(8.314 \text{ J K}^{-1} \text{ mol}^{-1})(273 \text{ K})} \right] \\ = 0.472 \text{ bar}$$

- ✓ 1.2 Calculate the second virial coefficient of hydrogen at 0 °C from the fact that the molar volumes at 50.7, 101.3, 202.6, and 303.9 bar are 0.4634, 0.2386, 0.1271, and 0.09004 L mol⁻¹, respectively.

SOLUTION

$$\frac{PV}{RT} = 1 + \frac{B}{V} + \frac{C}{V^2} + \dots$$

P/bar	50.7	101.3	202.6	303.9
V/L mol ⁻¹	0.4634	0.2386	0.1271	0.09004
PV/RT	1.035	1.064	1.134	1.205
(1/V)/mol L ⁻¹	2.158	4.191	7.868	11.106

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